

## Chemistry (New Book) - 9th Class Chemistry English Medium Chapter 8 Preparation

Q1. What is relationship between electropositivity and ionization energy.

**Ans 1:** Electropositivity depends upon the ionization energy which in turn depends upon size and nuclear charge of the atom. Small sized atoms with high nuclear charge have high ionization energy. Atoms having high ionization energy are less electropositive or metallic

Q2. Write down uses of Magnesium?

**Ans 1:** 1- Magnesium is used in flashlight bulbs and in fireworks  
2- It is used in the manufacture of light alloys

Q3. Why silver is not used in pure form?

**Ans 1:** Because the demand of silver for industrial purpose, etc that's why silver is not used in pure form

Q4. What do you mean by 24 carats of Gold?

**Ans 1:** Purity of gold is shown by carats that indicates the number of parts by weight of gold that is present in 24 parts of alloy

Q5. Why silver and gold are least reactive?

**Ans 1:** Silver and gold are least reactive because these metals do not lose their electrons easily. They do not have the tendency to make cations and due to their soft nature. They cannot be used as such.

Q6. Why copper is used for making electrical wire ?

**Ans 1:** Metals are malleable and ductile copper is a metal which is ductile and can conduct electricity easily, So copper is used for making electrical wires:

Q7. How oxygen reacts with group II A metals?

**Ans 1:** They are less reactive towards oxygen and oxides are formed on heating.

Q8. What are halogens?

**Ans 1:** In periodic table element of group -17 consists of fluorine, chlorine , bromine, iodine and astatine are called halogens

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Q9. Difference between Steel and Stainless steel?

**Ans 1:** Steel: Steel is an alloy of iron, manganese, cobalt, chromium and 4 percent carbon. Stainless steel. The best example of alloying is stainless steel. Which is good combination of iron, chromium and nickel

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Q10. Write any two chemical properties of metal.

**Ans 1:** 1-They have metallic bonding  
2-They usually form ionic compounds with non-metals

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