

Chemistry (New Book) - 9th Class Chemistry English Medium Chapter 7 Preparation

Q1. State the best method for protection of metal from corrosion.

Ans 1: The best method for protection against the corrosion of metals exposed to acidic conditions is coating the metal with other metal. Corrosion resistant metals like Zn, Sn and Cr are coated on the surface of iron to protect it from corrosion.

Q2. Define Electroplating?

Ans 1: Electroplating is depositing of one metal over the other by means of electrolysis.

Q3. What is meant by salt bridge?

Ans 1: Salt bridge is a U-shape and glass tube .
It consists of saturated solution of strong electrolyte supported in jelly type material

Q4. Define oxidation state?

Ans 1: "Oxidation state is the apparent charge assigned to an atom of an element in a molecule or ion.

Q5. Difference between Electrolyte and Galvanic cell?

Ans 1: Electrolyte
1- It consists of one complete cell, connected to a battery
2- Anode has positive charge while cathode has negative charge.

Ans 2: Galvanic:
1- It consist of two half cells connected through a salt bridge
2- Anode while cathode has positive charge.

Q6. Explain oxidation and Reduction Oxidation:

Ans 1:
Oxidation :
The process in which loss of electrons take place from an atom or ion is called as oxidation

Ans 2: Reduction :
Gain of electron is called reduction.

Q7. What is meant by Valency and Oxidation state?

Ans 1: Oxidation state is the apparent charge assigned to an atom of an element in a molecule in an ion. Valency the apparent charge of an atom, ion or molecules which are called Valency

Q8. Why steel is plated with nickel before the electroplating of chromium?

Ans 1: The steel is usually plated first with nickel on copper and then by chromium because it does not adhere well on the steel surface, Moreover, it allows moisture to pass through it and metal is stripped off. The nickel or copper provides adhesion and then chromium deposited over the adhesive layer of copper lasts longer. This type of electroplating resists corrosion and gives a bright silvery appearance to the object.

Q9. Difference between strong and weak electrolytes?

Ans 1: Strong Electrolytes.

The electrolytes which ionize completely in solution and produce more ions are called strong electrolytes

Ans 2: Weak electrolytes;

The substance which ionize to a small extent when dissolved in water and could not produce more ions are called weak electrolytes

Q10. How the half cells of a galvanic cell are connected? What is function of salt -bridge?

Ans 1: The two half cells of the galvanic cell are connected by salt bridge salt bridge is used to maintain the flow of ionic current and the electric neutrality.