Chemistry (New Book) - 9th Class Chemistry English Medium Chapter 2 Preparation

Q1. Why does the energy of electron increase as we move from first shell to second shell.

Ans 1: The energy of n electorn increases as we mvoe from the first to the second shell beacuse the second shell is farther awary from the nuclesus which means that electron in the second shell has high eneryg as compared to the electronin the first shell. This increase in distance from the nucleus results in a highe renergy level for electron in outer shell

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Q2. Why isotopes of an element show same chemicla properties while their physical properties are different.

Ans 1: Isotpes have same chemicla properties because they have same electronic configuration and have same atomic numebr. Isotopes have different physical properties beacues they have different mass number.

Q3. What are sub-atomic particles.

Ans 1: Sub -atomic particles are teh fundamnetal particles that make up atoms The three main sub-atomic particles are These particles were later named as electrons. It was also shown that electrons are teh subatmic particles of all elements.

Q4. Write application of radioactive isotopes.

Ans 1: Radioactive isotopes are used to test t he strength of metals and concreate mixture.
They are used tro generate cheap nuclear power
They also used to find oil fields.

Q5. How can atoms be seen.

Ans 1: The size of an atom is so small that it is not possible to see it with naked eye. However, a transmission electron microsocpe can be used to see atoms.

Q6. Why Gallium does not evaporate?

Ans 1: Gallium melting point is below body temperature so it is liquid at room temperature. it does not evaporate.

Q7. How many atoms does our body replace every year.

Ans 1: Every year our body replaces about 98% of its atoms.

Q8. What percentage of the atom's mass in contained in the nucleus.

Ans 1: Although the nucleus is less than ne hundred-thousand of this size of the atom it contains more than 99.9% of the mass of the atom.

Q9. What is gallium that make it unique from other mtals.

Ans 1: Gallium has many i ntersting properties. It melting point is below body temperatur eso it it liquid at room temperature . it has water like viscosity . it does not evaporate.

Q10. Defien relative isotopic masses.

Ans 1: An element usually consists of a few different isotopes with differentr mass numbers. These mass numbers are called relative isotopic massses. Each isotope will also have its own naturally occuring abundance which is called isotopic abundance.
