

Chemistry (New Book) - 9th Class Chemistry English Medium Chapter 1 Preparation

Q1. Why atomic mass of an atom is defined as relative atomic mass ?

Ans 1: Atomic mass of an atom is defined as a relative atomic mass , because atomic mass of an atom is the mass which compares or relates with 1/12th mass of carbon 12 isotope

Q2. Define chemical formula and give one examples .

Ans 1: Definition : A chemical formula is the symbolic representation the elements in the compounds and the ratio of the elements to one another

Ans 2: Example : Chemical formula of water is H₂O

Q3. Differentiate between biochemistry and industrial chemistry .

Ans 1: Biochemistry : Biochemistry is the branch of chemistry in which w study the structure , composition , and chemical reaction of substances found in living organisms

Ans 2: Industrial chemistry : Industrial chemistry is the branch of chemistry that deals with the manufacturing of chemical compounds in commercial scale is called industrial chemistry

Q4. Write down names of elements of first group and second period of the periodic table .

Ans 1: First group of periodic table consists of hydrogen , lithium , sodium , potassium , rubidium , cesium , francium

Ans 2: Second group of periodic table consists of lithium , beryllium , boron , carbon , nitrogen , oxygen , fluorine and ends at neon , a noble gas

Q5. How can we distinguish a metal from non-metal by a simple physical method ?

Ans 1: Metals : On heating metals usually become soft and flexible
They are difficult to cut

Ans 2: Non-Metals :Non metals are usually non-flexible on heating
They are easily cut or break

Q6. Define empirical formula and give examples.

Ans 1: It is the simplest whole number , ratio of atoms present in a compound." Empirical formula of benzene Empirical formula of Hydrogen peroxide

Q7. Write ant two chemical properties of non-metals .

Ans 1:

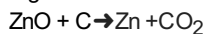
1 : Non-metals usually do not react with water

2: They do not react with dilute acids because non-metals are itself electron acceptors

Q8. Differentiate between oxidation and reduction.

Ans 1: Oxidation : Oxidation is defined as addition of oxygen or removal of hydrogen during a chemical reaction

e.g



Ans 2: Reduction :: Reduction is defined as addition of hydrogen or removal of oxygen during a chemical reaction

e.g $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}_2$

Q9. Define alloy and give example .

Ans 1: Alloy : alloy is a homogeneous mixture of one metal with one or more metals or non-metals

Ans 2: Example : The best example of alloy is 'stainless steel'

Q10. Name the most abundant non- metal in earth's crust?

Ans 1: Oxygen has has the highest percentage in earth's crust 47% and oceans 86% but second highest percentage in atmosphere 21%