

Chemistry - 10th Class Chemistry English Medium Chapter 15 Preparation

Q1. Give a use of kerosene oil

Ans 1: Used a domestic fuel , a special grade of it is used as jet fuel

Q2. How petroleum is extracted ?

Ans 1: Petroleum is extracted by drilling holes into earth's crust where the oil is formed

Q3. Write down the the names of four fractions obtained by the fractional distillation of residual oil ?

Ans 1: 1: Lubricant
2: Paraffin wax
3: Asphalt
4: Petroleum

Q4. Name preventive methods of water borne diseases.

Ans 1: Provision of safe water. Disposal of sewage Control of toxic chemicals.

Q5. How CO₂ prepared in the Solvay's process ?

Ans 1: CO₂ is prepared by heating limestone in a lime kiln . Then it is carried to carbonating tower

Ans 2: $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$

Q6. What role is played by pine oil in the forth floatation process ?

Ans 1: Oil coated or partials begins lighter come to the surface in the form of forth that can be skimmed

Q7. What is the principle of functional distillation ?

Ans 1: Refining is the process of separation of crude oil mixture into various useful products (functions) . It is carried out by a process called fractional distillation . The principle of fractional distillation is based upon separation depending upon their boiling points

Q8. Name two ores of copper

Ans 1: Copper glance Cu_2S
Chalcopyrite CuFeS_2

Q9. Define petroleum

Ans 1: Petroleum means rock oil . It is a complex mixture of several gaseous , liquid and solid hydrocarbons having water , salts and earth particles with it

Q10. Explain process of electro refining ?

Ans 1: Refining the impure metal by electrolysis is the most widely used process of refining metals . Example , electrolytic refining of copper is carried out in an electrolytic tank having copper sulphate solution in it . . Two electrodes : one of impure copper metal that acts as anode and the other of pure copper metal that acts as cathode are suspended in the electrolytic solution
On passing the electric current through the solution, anode (impure copper) dissolves to provide Cu^{2+} ions to the solution . These Cu^{2+} ions are discharged by gaining of electrons from the cathode . Thereby copper atoms deposit on the cathode , making it thick block of pure copper metal . The impurities like gold and silver settle down as anode mud
