Chemistry - 10th Class Chemistry English Medium Chapter 9 Preparation

Q1. why reversible reactions never complete?

Ans 1: As reversible reactions are those in which reactants combine to form product and product recombine to form the reactants that is why they never complete.

Q2. How the active mass is represented ?

Ans 1: Active mass is represented by square brackets { }

Q3. Write down the component of atmosphere and its percentage

Ans 1: The two major components of atmosphere are nitrogen and oxygen gases constitute 99% of the atmosphere

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Q4. What do you mean by equilibrium constant?

Ans 1: Ratio of the product of concentration of product raised to the power of the coefficient to the product of concentration of reactants raised to the power of co-efficient in a balance chemical equation.

Q5. Define chemical equilibrium state

Ans 1: When the rate of forward reaction take place at the rate of reverse reactions , the composition of the reaction mixture remain constant , it is called chemical equilibrium

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Q6. How dynamic equilibrium is established?

Ans 1: When the forward reaction become equal to the reverse reaction then equilibrium state is established Rate of forward reaction = Rate of Reverse reactions.

Q7. What is static equilibrium, explain with example.

Ans 1: When reaction ceases to proceed it is called static equilibrium. Example: A building remains standing rather than falling down because all of forces acting upon it are balanced.

Q8. Write down the use of oxygen .

Ans 1: Oxygen is used to prepare sulphur dioxide which is further used to manufacture king of chemicals sulphuric acid

Q9. Write down the use of Nitrogen.

Ans 1: These gases ae being used to manufacture chemicals since the advent of 20th century.

Q10. How dynamic equilibrium is established ?

Ans 1: When the forward reaction become equal to the reverse reaction then equilibrium state is established