

## Chemistry - 10th Class Chemistry English Medium Chapter 9 Preparation

Q1. Why amount of reactants and products do not change in reversible reaction ?

**Ans 1:** In dynamic equilibrium rate of forward reaction is always equal to rate of reverse reaction as in reversible reaction by dynamic equilibrium condition is appear that is why the amount will not change

Q2. What is direction of reversible reaction?

**Ans 1:** These reactions proceed in both direction. Forward direction Reverse direction.

Q3. How dynamic equilibrium is established ?

**Ans 1:** When the forward reaction become equal to the reverse reaction then equilibrium state is established

Q4. Which types of reactions do not go to completion ?

**Ans 1:** Reversible reaction never go to completion . Because in it products recombine to form the reactants

**Ans 2:** Characteristics :

- 1: They are represented by single arrow
- 2: They are supposed to be complete reaction

Q5. If a reaction has large value of  $K_c$  will it go to completion and why ?

**Ans 1:** The large values of  $K_c$  indicate that at equilibrium state reaction mixture consists of almost all product and reactants are negligible . It also shows that reaction has almost gone to completion reaction

Q6. What is meant by extent of reaction ?

**Ans 1:** It indicates how much reactants are converted into products and how far a reaction proceeds before reactant form the products

**Ans 2:**

Q7. How the active mass is represented ?

**Ans 1:** Active mass is represented by square brackets { }

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Q8. Why the reversible reactions do not go to completion ?

**Ans 1:** It is because in reversible reactions products re-combine to form the reactants in a same rate as the reactants from the product

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Q9. Why reversible reaction never complete ?

**Ans 1:** As reversible reactions are those in which reactions combine to form product and product recombine to form the reactants that us why they never complete

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Q10. Define chemical equilibrium state.

**Ans 1:** When The rate of forward reaction take place at the rate of reverse reaction, the composition of the reaction mixture remains constant, it is called chemical equilibrium.

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