

Chemistry - 10th Class Chemistry English Medium Chapter 9 Preparation

Q1. What is law of mass action ?

Ans 1: The rate at which a substance react is directly proportional to active mass and the rate of reaction is directly proportional to the product of the active mass of of the reacting substance

Q2. What do you mean by equilibrium constant?

Ans 1: Ratio of the product of concentration of product raised to the power of the coefficient to the product of concentration of reactants raised to the power of co-efficient in a balance chemical equation.

Q3. Why the reversible reactions do not go to completion ?

Ans 1: It is because in reversible reactions products re-combine to form the reactants in a same rate as the reactants from the product

Q4. Why at equilibrium state reaction does not stop ?

Ans 1: It is because the forward reaction take place at the same reaction as of reverse but in opposite direction . So reactants convert into product and products into reactants

Q5. Define chemical equilibrium state.

Ans 1: When The rate of forward reaction take place at the rate of reverse reaction, the composition of the reaction mixture remains constant, it is called chemical equilibrium.

Q6. Give the characteristics of reversible reaction

Ans 1: The reactions in which products recombine to product the reactants are called reversible reactions

Ans 2: Presentation :

1: Forward direction

2: Reverse direction

Ans 3: Completion :

They never go to completion

Q7. What are irreversible reaction?

Ans 1: The reactions in which the product do not recombine to form the reactants.

Q8. Write down the use of Nitrogen.

Ans 1: These gases are being used to manufacture chemicals since the advent of 20th century.

Q9. What are the characteristics of a reaction that established equilibrium state at once?

Ans 1: If a reaction has very small value of K_c will attain the equilibrium state at once

Q10. Why amount of reactants and products do not change in reversible reaction ?

Ans 1: In dynamic equilibrium rate of forward reaction is always equal to rate of reverse reaction as in reversible reaction by dynamic equilibrium condition is appear that is why the amount will not change
