

## PPSC Physics Chapter 3 Thermal Properties of Matter

Sr	Questions	Answers Choice
1	What is a thermal properly of a material that determines the quantity of energy required to change the phase of a unit mass of that substance.	A. Specific heat B. Latent heat C. Internal energy D. Specific energy
2	How does heat transfer between objects.	A. From cold to hot objects     B. From hot to cold objects     C. By electromagnetic radiations     D. From hotter to hottest objects.
3	Which thermometer is called sprint thermometer	A. Alcohol thermometer B. Mercury in glass thermometer C. Gas thermometer D. Radiation thermometer
4	On what factor the internal energy of a thermodynamic system depend upon.	A. History B. State C. Process D. Surroundings
5	A sample of an ideal gas may i) energy adiabatically, or ii) Expand isothermally. the net flow of heat into the gas from the exterior is.	A. Positive is each case B. Negative for i) and positive for ii) C. Zero for i) and positive for ii) D. Positive for i) and negative for ii)
6	A given quantity of an ideal gas is at pressure P and temperature T What is the isothermal bulk modulus of the gas.	A. 2/3 P B. P C. 2P D. 3/2 P
7	In onw cycle of a steam engine there are how many dead centres.	A. 1 B. 2 C. 3 D. zero
8	In an isochoric process.	A. Volume changes B. Pressure changes C. Temperature changes D. Volume remains constant
9	Under steady state the temperature of a body	A. Increase's with time B. Decreases with time C. Does not charge with time D. None of the above
10	What happen to entropy in an irreversible cycle.	A. No gain in entropy B. No change in entropy C. Loss of entropy D. A net gain of entropy
11	The expression PV/KT represents.	A. Number of moles of the gas B. Number of molecules in the gas C. Total mass of the gas D. Density of the gas
12	The volume of a goas at constant pressure is directly proportional to the temperature as measured on the.	A. Celsius scale B. Kelvin scale C. Fahrenheit scale D. Baume scale
13	When ever a system is made to complete a cyclic process the work done during the complete cycle.	A. Is zero B. Is negative C. Is positive D. Depends upon the path followed
14	Which law states that if two systems are in thermal equilibrium with a third system they are also in the thermal equilibrium with each other.	A. Third law of thermodynamics     B. Second law of thermodynamics     C. First law of thermodynamics     D. Zeroth law of thermodynamics
15	What is the total entropy change during an reversible cycle.	A. Unity B. Infinite C. Zero D. Cannot be detected

16	The SI unit of heat capacity is.	A. kg J B. Kg J-1 C. J K-1 D. Kg J-1 K-1
17	When a perfect gas is supposed to expand freely in an isolated vessel the gas has undergone.	A. An increase in pressure B. An increase in temperature C. A change in phase D. A change in entropy
18	Which of the following is the science of measuring the heat of chemical reactions or physical changes.	A. Thermometry B. Calorimetry C. Telemetry D. Photometry
19	How solid hydrogen is obtained.	A. By cascade process     B. By joule kelvin effect     C. By adiabatic expansion     D. Lowering temperature below melting point
20	What is the heat required in Kilo joules when the temperature of 100 g of copper is raised thgough 20 K/. Specific heat capacity of copper is 0.4 x 10-3 kg -1 k-1	A. 0.4 kj B. 0.8 kj C. 400 kj D. 800 kj