

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	During the last two centuries, the atmospheric CO ₂ contents are increased by	A. 15% B. 25% C. 35% D. 50%
2	Which of the following trihalides of nitrogen behaves as the weakest base.	A. NF ₃ B. NCl ₃ C. NBr ₃ D. NI ₃
3	The basic strength of hydrides of group 15 elements vary in the following order.	A. NH ₃ > PH ₃ > AsH ₃ > SbH ₃ > BiH ₃ B. PH ₃ > NH ₃ > AsH ₃ > SbH ₃ > BiH ₃ C. BiH ₃ > NH ₃ > PH ₃ > AsH ₃ > SbH ₃ D. NH ₃ > PH ₃ > SbH ₃ > AsH ₃ > BiH ₃
4	Aluminothermy used for on the spot welding of large iron structures is based upon the fact that.	A. As compared to iron, aluminium has greatest affinity for oxygen. B. As compared to aluminium, iron has greater affinity for oxygen C. Reaction between aluminium and oxygen endothermic D. Reaction between iron and oxygen is endothermic
5	On the basic of CFT the bonding between the metal and ligand is totally	A. Ionic B. Covalent C. Coordinate D. Metallic
6	Nitric acid is used in manufacturing of.	A. Explosive B. H ₂ SO ₄ C. Fertilizer D. All above
7	The atomic number of Potassium is 19 and that of manganese is 25. Although the coloured of MnO ₄ is dark violet yet the K ⁺ is colourless.. This is due to the fact that	A. Mn is a transition element while K ⁺ is not B. [MnO ₄] ⁻ is negatively charged while K ⁺ has a positive charge C. The effective atomic number of Mn is [MnO ₄] ⁻ is 26 while for K ⁺ the atomic number is 18 D. The Mn in a high positive oxidation state allows charge transfer transitions
8	The common oxidation state of lanthanides is.	A. +3 B. +2 C. +1 D. +4
9	Has maximum property of catenation.	A. C B. Si C. Sn D. Pb
10	Zeigler Natta catalyst is.	A. Pt/PtO B. TiCl ₄ /Al(C ₂ H ₅) ₃ C. Pt/Rh D. Pt
11	Potassium crystallizes in a body centered lattice. Hence, the coordination number of potassium in potassium metal is.	A. 4 B. 6 C. 8 D. 12
12	In vinyl cyanide, the number of a bonds in	A. 2 B. 3 C. 1 D. 4
		A. 180 ^o

13	The bond along Sp ² hybridization is.	B. 120 ^o C. 109.5 ^o D. 160 ^o
14	Which one of the following ions is colourless.	A. Cu ⁺ B. Ni ²⁺ C. Co ²⁺ D. Fe ³⁺
15	Bromine is soluble in	A. Alcohol B. Water C. Chloroform D. All above
16	Oxalic acid when heated with conc. H ₂ SO ₄ it gives out.	A. H ₂ O and CO ₂ B. CO and CO ₂ C. CO ₂ and H ₂ S D. Oxalic sulphate
17	In order to understand the nature of H, bond the theory has been suggested.	A. Electrostatic approach B. Molecular orbital approach C. Valence bond approach D. All the above approaches
18	Which of the following is not a proper use of Ni.	A. It is used as catalyst B. It is used in alloy formation C. It is used in the preparation of Monel metal D. It is attached by alkalis
19	Hemimorphite is an example of.	A. Orthosilicate B. Pyrosilicate C. Cyclic silicate D. Meta silicate
20	What types of bonding occurs in d-block elements.	A. Ionic B. Covalent C. Metallic D. Both B and C