PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	The ionization potential of K would be numerically equal to.	A. Electron affinity of Ar B. Electromagnetically of K C. Electron affinity of K+ D. ionization energy of Ca
2	Which ionization Potential in the following equations involves the greatest amount of energy.	A. Na = Na+ + e B. K = K+ +e C. C2+ = C3+ +e D. Ca+ = Ca2+ + e
3	The decreasing order of the second ionization energies of K, Ca and Ba is	A. K > Ca > Ba B. Ca > Ba > K C. Ba > K > Ca D. K > Ba > Ca
4	Which of the following elements would have the lowest first ionization energy	A. Mg B. Rb C. Li D. Ca
5	In group 17, the element with highest first ionization enthalpy belongs to.	A. Period 1 B. Period 2 C. Period 7 D. Period 6
6	In whihc period, the element with least ionization enthalpy belong to	A. Group I B. Group 2 C. Group 17 D. Group 18
7	Which of the following represents the correct order of ionic radii	A. La+ < Na+ < K+ < Rb+ B. Li+ > Na + > K+ >Rb+ C. Li = Na + = K+ = Rb+ D. Rb +> Na+> K+ > Li+
8	Which of the following ions does not have the electronic configuration same as that of neon.	A. F- B. O2- C. Na+ D. Ca2+
9	The size of iso electronic species - F- , Ne, and Na+ is affected by	A. Nuclear charge (Z) B. Valence principal quantum number (n) C. Electron electron interaction in the outer orbital D. None of the factors because their size to the same.
10	Which of the following ions is smallest in size.	A. F- B. Cl- C. I- D. Br-
11	The ions Sc3+ , Ca2+ and K+ have same electronic configuration as that of.	A. Neon B. Argon C. Krypton D. Xenon
12	Atomic volume of C, N, O and F are in the order	A. C > N > F > O B. C > N > O > F C. F > O > N >C D. N > C > O >F
13	The correct order of ionic radii for the following ions is.	A. S²- < P³⁻ < Cl- < K+ B. Cl- > S2+ > P3- > k+ C. K+ > Cr > S2+ > P3- D. P3+ > S2 > Cl - > K+
14	In graph of atomic volume versus atomic weight the elements corresponding to peaks in the curve belong to.	A. Group 1 B. Group 18 C. Group 4 D. Group 14

15	An element with atomic number 20 is placed in which period of the periodic table.	A. 1 B. 2 C. 3 D. 4
16	Which group contains elements that exist as monoatomic molecules.	A. 1 B. 2 C. 14 D. 18
17	Elements in the same vertical group of the periodical have same	A. Number of electron B. Atomic number C. Number of valence elections D. Electronic configuration
18	In the long form of periodic table, elements are arranged according to.	A. Increasing atomic number B. Decreasing atomic number C. Increasing atomic mass D. Decreasing atomic mass
19	Lother Meyer plotted a graph showing variation of.	A. Atomic volume with increase in atomic number B. Atomic volume with increase in atomic weight C. Atomic redii with increase in atomic weight. D. Atomic weight which increase in atomic number
20	The law of trinds was proposed by	A. Dobereiner B. Newlands C. Lother Mayer D. Chancourtois