PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	The correct order of electron affinities is.	A. C > Si, > Na > Ar B. Si > Cl > Na > Ar C. C > Na > Si > Ar D. C > Si > Ar > Na
2	The correct order of electron affinities of Sl, P, and Cl is.	A. P > Si > Cl B. Cl > P > Si C. Cl > Si, > P D. Si > P . Cl
3	The magnitude of electron affinity depends on.	A. Atomic size B. Nuclear charge C. Electronic configuration D. All of the above
4	The elements with highest electron affinity belongs to.	A. Period 2 , group 17 B. Period 3, group 17 C. Period 2, group 18 D. Period, 2 , group 1
5	In a period, the element with biggest electron affinity belong to.	A. Group 1 B. Group 2 C. Group 17 D. Group 18
6	The first ionization energy of Mg is lower than	A. Na B. Ca C. Al D. Be
7	Which of the following elements has the highest ionization energy.	A. Na B. Si C. Ar D. Cl
8	Which one of the following statement is incorrect in relation to ionization enthalpy.	A. Ionization enthalpy increase for each successive electron B. The greatest increase in ionization enthaly is experienced on removal of electron from core noble gas configuration C. End of the valence electron is marked by a big jump in ionization enthalpy D. Removal of electron from orbitals bearing lower value to easier than from orbital having highest n value.
9	Which of the following elements has the highest third ionization energy.	A. Sodium B. Magnesium C. Aluminum D. Silicon
10	Among the elements A,B,C and D having atomic numbers 9,10,11, and 12 respectively, the correct order of ionization energies is.	A. A > B > C >D B. B > A > D > C C. B > A > C > D D. D > C > B > A
11	Which of the following statement is not correct.	A. The element with highest IE belongs to group 18 B. In each period the element with lowest IE belongs to group I C. In each period the element with highest IE is a noble gas D. In the second period as we move from left to right, ionization energy increases regularly.
12	With which one of the following configurations, the lowest value of first IE is associated.	A. ls2, 2s2, 2p6, 3s1 B. 1s2, 2s2, 2p5 C. 1s2, 2s2, 2p6 D. 1s2, 2s2, 2p6, 3s2, 3p2
13	Which of the following configuration is associated with biggest jump between second and	A. 1s2, 2s2, 2p2 B. 1s2, 2s2, 2p6, 3s1

	third IE.	C. 1s2,2s2,2p6, 3s2 D. 1s2, 2s2,2p6
14	The element with the highest first ionization potential is.	A. Boron B. Carbon C. Nitrogen D. Oxygen
15	The correct order of second ionization potential of carbon nitrogen, oxygen and fluorine is.	A. C > N > O > F B. O > F > N > C C. O > N > F > C D. F > O > N > C
16	The first ionization energy in electron volts of nitrogen and oxygen atoms are respectively given by.	A. 14.6, 13.6 B. 13.6, 14.6 C. 13.6, 13.6 D. 14.6, 14.6
17	The process requirieng the absorption of energy of.	A. F = F B. Cl = Cl C. H = H D. O = O
18	The ionization energy of N is more than that of oxygen becasue.	A. Nitrogen has half filled p orbitals B. Nitrogen atom is smaller in size than oxygen atom C. Nitrogen contains less number of electrons D. Nitrogen is less electronegative
19	Which of the following iso -electronic ion would require least energy for the removal of electron.	A. Ca2+ B. Cl- C. Ar- D. K+
20	Which of the following iso electronic spices has the highest IE.	A. Ne B. Na+ C. F D. O2-