PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Flourine differs from the other members of its own group due to.	A. Its small size and low bond energy B. Its higher electornegativity C. None-availability of d-orbitals in its valence shell D. All the above
2	The Lewis structure of which of the following does not have coordinate bond.	A. SO2 B. HNO3 C. H2SO4 D. HNO2
3	The three isotopes of hydrogen differ from one another in	A. Atomic number B. Number of protons C. Nuclear charge D. Nuclear mass
4	The aluminium alloy used to make parts of aircrafts is.	A. Magnalium B. Aluminium bronze C. Duralumin D. All of the these
5	The ionization potential of K would be numerically equal to.	A. Electron affinity of Ar B. Electromagnetically of K C. Electron affinity of K+ D. ionization energy of Ca
6	Pick out the incorrect statement for SO2	A. It turns filter paper moistened with acidified K2Cr2O7 B. It turns starch iodate paper blue C. It does not react with chlorine in presence of charcoal D. It decolourises acidified KMnO4 solution.
7	Which of the following molecule contains two dative bonds according to Lewis structure.	A. NH3 B. SO3 C. PCl5 D. BF3
8	How many unpaired electron are there in a strong field iron (II) octahederal compled.	A. 0 B. 1 C. 2 D. 4
9	In the formation of H2O molecule, the oxygen atom makes use of.	A. 2p orbitals B. sp hybrid orbitals C. Sp2 hybrid orbitals D. Sp3 hybrid orbitals
10	H2SO4 is used	A. In the preparation of aqua regia B. In the purification of gold and silver C. In the dental filling D. None of above
11	In whihc period, the element with least ionization enthalpy belong to	A. Group I B. Group 2 C. Group 17 D. Group 18
12	The compound which does not act as Lewis acid is.	A. BF3 B. AlCl3 C. BeCl2 D. SnCl4
13	The most stable oxidation state of chromium is.	A. +6 B. +3 C. +4 D. +2
14	Which of the following elements has the highest value of second ionization energy.	A. Lithum B. Beryllium C. Boron D. Magnesium

15	The carbonate of which of the following will have highest lattice energy.	A. Barium B. Magnesium C. Calcium D. Strontium
16	The atomic number of potassium is 19 and that of mangness is 25 Although the colour of MnO4 is dark violet yet the K+ is colourless this is due to the fact that.	A. Mn is a transition element while K+ is not B. [MnO4]- is negatively charged while K+ has positive charge C. The effective atomic number of Mn is [MnO4]- is 26: while for K+ the atomic number is 18 D. The Mn is a high positive oxidation state allows charge transfer transitions.
17	The element having electronic configuration 1s2, 2s2, 3s2, 3p3 is.	A. Trivalent only B. Tetravalent only C. Trivalent and pentavalent D. Pentavalent only
18	Among group IA elements, melting point	A. Increases down the group B. Decreases down the group C. Do not show any regular trend D. Remains constant
19	The solution of the transition metal complexes having one or more unpaired electrons in the d-orbital are.	A. Coloured B. Colourless C. White D. None of above
20	Example of peseudohalonge group.	A. Cyanogen B. Thiocyanogen C. Selenocyanogen D. All above