

## PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Among sodium phosphate, sodium sulphate and sodium chloride the solubility in water increases as.	A. Chloride &gt; Phosphate &gt; Sulphate B. Sulphate &gt; Pohosphate&gt; Chloride C. Chloride &gt; Sulphate &gt; Phosphate D. Phosphate &gt; Chloride &gt; Sulphate
2	The electronic configuration of sodium (Z=11)	A. 1s2, 2s2, 2p4 B. 1s2, 2s2, 2p6, 3s2, 2p5 C. 1s2, 2s2, 2p6, 3s1 D. 1s2, 2s2, 2p6, 3s2
3	Which among the following hydride is ionic in nature.	A. Ammonia B. Protium oxide C. Calcium hydride D. Sulphane
4	Ionic compounds in general possess both	A. High melting point and non - directional bonds B. High melting points and low boiling points C. Directional bonds and low boiling points D. High solubility in polar and non - polar bonds.
5	Solid sodium chloride does not conduct electricity be cause.	A. In solid NaCl, no ions are present B. Solid NaCl is covalent in nature C. In solid NaCl, there is no mobility of ions D. In solid NaCl, there are no electrons.
6	Ionic reactions mainly take place in.	A. Aqueous solutions and organic solvents of high polarity B. Non aqueous solvents of low polarity C. Gaseous state D. Solid state
7	Which of the following halide has lowest melting point.	A. NaCl B. NaF C. NaBr D. NaI
8	Which of the following has the highest melting point.	A. NaCl B. KCl C. MgO D. BaO
9	Among the solvents given below, with dielectric constant (E) given in parentheses which has highest solubility of KCl?	A. Benzene (E=0) B. Carbon disulphide (E = 0) C. Methanol (E =32) D. Acetone (E = 2)
10	Which of the following will exhibit variable electro Valency due to inert pair effect.	A. Fe B. Sn C. K D. Both Fe and Sn
11	Which compound among the following does not contain an ionic bond.	A. NaOH B. HCl C. KaS D. LiH
12	Which of the following compounds is electrovalent in nature.	A. SO2 B. ICl C. KBr D. CHI3
13	Which of the following compound does not following octet rule.	A. CS2 B. PBr3 C. IBr D. BrF3

14	An ionic compound $X + Y^-$ is most likely to be formed if	A. Ionization enthalpy of X is high electron gain enthalpy of Y is low B. Ionization enthalpy of X is high electron gain enthalpy of Y is high C. Ionization enthalpy of X is low, electron gain enthalpy of Y is low D. Ionization enthalpy of X is low electron gain enthalpy of Y is high
15	Four elements A, B, C, D have atomic numbers $Z, Z+1, Z+2$ and $Z+3$ respectively. If $Z$ is 9, then bond between which pair of elements will be ionic.	A. A and C B. D and C C. D and B D. B and C
16	Which of the following orbitals has maximum penetration effect.	A. s B. p C. d D. f
17	Among the elements of third period, the element with lowest boiling point belongs to group.	A. 1 B. 14 C. 16 D. 18
18	Among the elements of second period the element with highest melting point belongs to group.	A. 1 B. 14 C. 17 D. 18
19	Which of the following has the highest melting point.	A. NaCl B. LiCl C. KCl D. RbCl
20	Considering the elements F, Cl, O and N, the correct order of their chemical reactivity in terms of oxidizing property is.	A. $F > Cl > N$ B. $F > O > Cl > N$ C. $Cl > F > O > N$ D. $O > F > N > Cl$