PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	A stable molecule is a group of atoms held together by	A. Chemical forces B. Physical forces C. Valence force D. None of above
2	The forces which holds the atoms together in a molecule is called	A. Ionic bond B. Covalent bond C. Co ordinate bond D. Chemical bond
3	A molecule MX4 has a square planar shape, The number of non bonding pairs of electrons around M is .	A. 2 B. 1 C. 0 D. 3
4	Which of the following molecule does not contain the covalent bond between similar atoms.	A. N2H4 B. F2O2 C. H2F2 D. H2O2
5	In the formation of H2O molecule, the oxygen atom makes use of.	A. 2p orbitals B. sp hybrid orbitals C. Sp2 hybrid orbitals D. Sp3 hybrid orbitals
6	The geometry of the molecule is primarily decided by	A. Bond pairs around the central atom B. No of k bond around the central atom C. No of bond pairs as well as lone pairs around the central atom D. No. of lone pairs on central atom
7	The configuration of valence shell of certain atom X is 3s2 , 3p5, which valences can it exhibit.	A. 1,3 only B. 1,5 only C. 1,3,5,7 D. 1,3,4
8	In vinyl cyanide, the number of a bonds in	A. 2 B. 3 C. 1 D. 4
9	In which of the following species the bonds are non directional.	A. NCl3 B. RbCl C. BeCl2 D. BCl3
10	Which of the following statements is correct.	A. A sigma bond is weaker than a pi pond B. There are four coordinate bonds in the Lewis structure of NH4+ ion. C. The 1 covalent bond is directional in nature D. A single bond between the two atoms cannot be re bond.
11	The Lewis structure of which of the following does not have coordinate bond.	A. SO2 B. HNO3 C. H2SO4 D. HNO2
12	Which of the following molecule contains two dative bonds according to Lewis structure.	A. NH3 B. SO3 C. PCl5 D. BF3
13	The number of electrons involved in bonding in Lewis structure of oxalate ion is	A. 20 B. 14 C. 22 D. 18
14	In the Lewis structure of H2SO4 molecule the total number of unshared electrons in valence shell of various atoms is.	A. 8 B. 16 C. 12

		D. 20
15	Which of the following statements in incorrect.	A. Sodium hydride is ionic B. Beryllium chloride is covalent C. CCl_4 gives a white ppt with $AgNO_3$ solutions. D. Bonds in NaCl are non directional
16	Which of the following is not a characteristic of covalent compound.	A. They have low melting and boiling points. B. They ionize on dissolution in polar solvents C. Their molecules have definite geometry D. They are generally insoluble in water
17	In bi sulphate ion, the formal charge on sulphru atom is.	A. +1 B. +2 C. +4 D. +6
18	Which of the following element has six electrons in the valance shell but cannot exhibit a maximum co valency of six.	A. Sulpher B. Oxygen C. Salenium D. Both A and B
19	Which element out of the following can exhibit a maximum co valency of seven.	A. Chlorine B. Fluorine C. Sulphur D. Both Cl and F
20	The Lewis formula of $SOCl_2$, the total number of bond pairs and lone pairs of electron around sulphur are.	A. 2 ,1 B. 2,2 C. 3,1 D. 3,0