

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Molecules have zero dipole moment	A. CO ₂ B. BCl ₃ C. CH ₄ & CCl ₄ D. All above
2	The polarity of bonds can lead to polarity of molecules and affect	A. Melting point B. Boiling point C. Solubility D. All of above
3	The bond length is measured by	A. X-ray diffraction B. Neutron diffraction C. Microwave spectroscopy D. All of above
4	The types of coordinate compounds.	A. Labila B. Inert C. Both A and B D. None of above
5	Co ordinate compounds are	A. Polar B. Non polar C. Dem polar D. None of above
6	Al ₂ Cl ₆ is an example of	A. Ionic bond B. Covalent bond C. Co ordinate bond D. Metallic bond
7	The acetylene molecule contain a	A. Single bond B. Double bond C. Triple bond D. Co ordinate bond
8	A type of a chemical bond which is formed by the mutual sharing of electrons between combining atoms of the same or different elements is called.	A. Ionic bond B. Covalent bond C. Co ordinate Covalent bond D. Metallic bond
9	Titanium dioxide shows the lattice strcuture.	A. Filluorite B. Rutile C. Wurtzite D. Zeolite
10	The coordination number of closely packed hexagonal is.	A. 4 B. 6 C. 8 D. 12
11	Mostly used solvents for ionic compounds.	A. Liquid ammonia B. Liquid SO ₂ C. Liquid HF D. All above
12	By applying an external force the ionic solid can be easily broken to powder form so the ionic solid are highly	A. Hard B. Brittle C. Tough D. Soft
13	The unit of sodium chloride structure is.	A. Linear B. Cubic C. Tetrahedral D. Square planner
14	Solid substances consist of an ordered any of ions and solid as a whole is electrically.	A. Conductor B. Neurtal C. Acidic D. Basic
15	Variable electrovalency is due to the following reasons.	A. Unstable configuration of core B. Inset electron pair effect C. All of above D. None of above

16	Stable metal ions structures are.	A. Noble gas structure B. Is electron group structure C. Transition metal in structure D. All of the above
17	The bond formed by complete transfer of electrons from electropositive to more electronegative atom is called.	A. Ionic bond B. Covalent bond C. Metallic bond D. Coordinates bond
18	Non localised bonds are referred as	A. Metallic bond B. Long range bonds C. Ionic bond D. Covalent bonds
19	Ionic bond are also forces called as.	A. Polar bond B. Electrovalent bond C. None polar bond D. Both A and B
20	When two H atoms approach each other then forces operates.	A. Attractive forces B. Repulsive forces C. Attractive and repulsive D. None of above