PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Which among the following is insoluble in water.	A. LiOH B. KOH C. NaOH D. RbOH
2	Example of intra molecular hydrogen bonding.	A. O-nitrophenol B. O-hydroxy benzaldehyde C. O- hydroxy benzoic acid D. All of the above
3	Hybridization involves.	A. Orbitals of same atom with slightly different energies. B. Orbitals of different atoms, but with equal energies. C. Orbitals of the same atom but with widely different energies. D. Orbitals of different atoms with different energies.
4	Which element out of the following can exhibit a maximum co valency of seven.	A. Chlorine B. Fluorine C. Sulphur D. Both Cl and F
5	Nitric acid has the property	A. Nitrating B. Reducing C. Redoxing D. None of above
6	BCl3 is a planar molecule because B atom is.	A. sp2 hybridized B. Sp3 hybridized C. sp hybridzed D. sp3 d hybridized
7	Which of the following orbitals has maximum penetration effect.	A. s B. p C. d D. f
8	The alkali metal with highest melting point is	A. K B. Na C. Li D. Ca
9	The wire of flash bulb is made up of.	A. Cu B. Ag C. Mg D. Ba
10	Which element out of the following can ehibit a maximum con valency of seven.	A. Chlorine B. Sulphur C. Fluorine D. both Cl and F
11	The H2SO4 obtained by the contact process having purity	A. 70% B. 74% C. 78% D. 82%
12	Ca H2 on reaction with water liberates	A. H2 B. O2 C. Botha of these D. None of these
13	Of the following the commonly used n the laboratory desiccator is.	A. Anhyd. Na2Co3 B. Anhyd Ca Cl2 C. Dry NaCl D. None of the above
14	A thionic acid	A. H2S2O3 B. H2S2O6 C. H2S2O8 D. H2S2O7

15	The type of bonding in HCl is	A. Pure covalent B. Polar covalent C. Highly polar D. Hydrogen bonding
16	Potassium reacts with excess of oxygen to form	A. K2O B. K2O2 C. KO2 D. K2O3
17	Chlorine is used in	A. Sterilizatin of water B. Extractionof gold C. Bleaching of cotton D. All above
18	The most electronegative element of the third period is.	A. F B. P C. Br D. Cl
19	Flourine differs from the other members of its own group due to.	A. Its small size and low bond energy B. Its higher electornegativity C. None-availability of d-orbitals in its valence shell D. All the above
20	Which show maximum number of oxidation states in 3d series.	A. Mn B. Ni C. Co D. Zn