

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Metals are	A. Transparent B. Translucent C. Opaque D. None of above
2	Metal are generally elements	A. Electronegative B. Electropositive C. Neutral D. None of the above
3	Metal crystallize is system having co ordination number	A. 8 B. 12 C. 14 D. any one of above
4	H-Bonding also ox in ling system like	A. Protein B. DNA C. Both A and B D. None of above
5	H-Bond has a preferred bonding direction like	A. Ionic bond B. Covalent bond C. Co ordinate bond D. None of these
6	H-Bond has more energy than the van der Waals forces i.e.	A. 1.0 kcal/mole B. 2.0 kcal/mole C. 10.0 kcal/mole D. 20.0 kcal mole
7	Excluding H-atom , Hydrogen bond never involves more than atoms.	A. One B. Two C. Three D. Four
8	Hydrogen bond is not electrostatic in nature is stated by	A. Electrostatic approach B. Valence bond approach C. Molecular orbital approach D. None of the above
9	In order to understand the nature of H , bond the theory has been suggested.	A. Electrostatic approach B. Molecular orbital approach C. Valance bond approach D. All the above approaches
10	Example of inter molecular H-bonding is	A. NH ₃ and H ₂ O B. HF C. CH ₃ COOH D. All of above
11	Example of intra molecular hydrogen bonding.	A. O-nitrophenol B. O-hydroxy benzaldehyde C. O- hydroxy benzoic acid D. All of the above
12	The bond order for BO molecule is.	A. 2.5 B. 3.0 C. 2.0 D. 3.5
13	The bond order gives the following valuable information.	A. Stability of the molecules of ions B. Bond dissociation energy and bond length C. Magnetic properties D. All of the above
14	d ² sp ³ is oriented in a manner	A. Trigonal B. Tetrahedral C. Octahedral D. Trigonal bipyramidal
15	PCl ₅ is an example of hybridization	A. d sp ³ B. d ² sp ² C. sp ² D. sp ³

		D. sp^3
16	The bond angle along sp^2 hybridization is.	A. 180° B. 120° C. 109.5° D. 160°
17	BCl_3 is an example of hybridization	A. sp B. sp^2 C. sp^3 D. None of above
18	A covalent bond which is formed between two atoms by the overlap of atomic orbitals along their axis is called.	A. Pi bond B. Sigma bond C. Polar bond D. Non polar bond
19	Valence bond theory is also called as	A. Electron pair theory B. Band theory C. Electron gas theory D. Electron pool theory
20	Pauling has suggested that the calculation of energy can be improved by considering.	A. Screening effect B. Polarization effect C. Both A and B D. None of above