

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	In which of the following compounds does hydrogen bonding occur.	A. CCl ₄ B. NaH C. HI D. NH ₃
2	NH ₃ has a not dipole moment while BF ₃ has zero dipole moment Thsi is because.	A. NH ₃ is not a planar molecule while BF ₃ is a planar molecule. B. NH ₃ is a planar molecule, while BF ₃ is a planner molecule. C. Fluorine is more electronetative than nitrogen D. Born is more electronegative than nitrogen
3	The element having electronic configuration 1s ² , 2s ² , 3s ² , 3p ³ is.	A. Trivalent only B. Tetravalent only C. Trivalent and pentavalent D. Pentavalent only
4	The percentage of s-character in the hybrid orbitals sp, sp ² and sp ³ follows the pattern.	A. sp ³ > sp ² > sp B. sp > sp ² > sp ³ C. sp = sp ² > sp ³ D. sp = sp ² = sp ³
5	The state of hybridization of carbon in CO ₂ is	A. sp ² B. sp C. sp ³ D. dsp ²
6	Which one of the following does not exhibit paramegnetion.	A. NO B. NO ₂ C. ClO ₂ D. ClO ₂ ⁻
7	Which of the following ahs non zero dipole moment.	A. NH ₃ B. SF ₆ C. BF ₃ D. CO ₂
8	The type of bonding in HCl is	A. Pure covalent B. Polar covalent C. Highly polar D. Hydrogen bonding
9	Which one has a co ordinate bond.	A. Al ₂ Cl ₆ B. BF ₃ C. NaCl D. O ₂
10	Which of the following proportion is associated with the covalent nature of the compound.	A. It conducts electricity in molten stater or aqueous state B. It is a non electrolyte C. It has high m.p. D. It is a compound of a metal and non metal.
11	The unequal sharing of bonded pair of electrons between the two atoms in a molecule causes.	A. Dipole B. Radical formation C. Decomposition of found D. Covalent found
12	CCl ₄ has zero dipole moment because of.	A. Planar structure B. Tetrahedral structure C. Similar size of C and Cl atoms D. Similar electrons affinity of C and Cl
13	The important condition for the formation of chemical bond is that.	A. Their electron clouds should not diffuse B. Both atoms should have high electron affinities. C. Both atoms should have same electronegativities D. The process should be accompanied by the lowering in

potential energy.

14	Example of linear geometry	A. XeF ₂ B. F ₂ and HgCl ₂ C. CdI ₂ AND Ag Cl ₃ D. All of the above
15	The gap between occupied and the unoccupied orbitals is not very large and the conduction of electricity is negligible at lower temperature and appreciable at high temperatures then it will be.	A. Good conductor B. Non conductor C. Semi conductor D. None of the above
16	The energy gap between two bands so large that it effectively prevents the promotion of electron from the lower to the higher band such energy gap all called.	A. Ionization zone B. Dissociation zone C. Distinction zone D. Forbidden zone
17	A combination of atomic orbitals produces a large number of closely spaced energy states known as.	A. Packet of energy B. Band of energy C. Both a and b D. None of the above
18	Metallic bond is treated essentially as in character	A. Ionic B. Covalent C. Polar D. Non polar
19	Electron gas theory fails to explain	A. Specific heat of metals B. Electrical and thermal conductivity C. Paramagnetic behavior of metals D. All of the above
20	Electron gas theory is able to explain	A. Metallic lusture and optical properties B. Malleability and ductility C. High electrical and thermal conductivity D. All of the above