

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Which among the following is least soluble in water.	A. NaF B. LiF C. KF
		D. CsF A. Barium
2	The carbonate of which of the following will have highest lattice energy.	B. Magnesium C. Calcium D. Strontium
3	In which of the following characteristics does hydrogen resemble halogens.	A. Hydrogen is the lightest gas B. H atoms contains one electron each C. Hydrogen forms ionic hydrides with alkali metals D. Hydrogen has three isotopes.
4	Which of the following generally increases on going from top to bottom in a group.	A. Metallic charcter B. Electronegativity C. Oxidising behaviour D. Raducing behaviour
5	Pick out the incorrect statement about K2Cr2O7	A. It is thermally stable B. It dissolves in alkali to form chromate C. It oxidizes acidified FeSO4 solution to Fe2(SO4)3 D. It is used as cleansing agent for glassware. etc. When mixed with cold con. H2SO4
6	All the halogen form oxyacide, except	A. Florine B. Chlorine C. Bromine D. lodine
7	A molecule MX4 has a square planar shape, The number of non bonding pairs of electrons around M is .	A. 2 B. 1 C. 0 D. 3
8	Perdisulphuric acid is.	A. Marshal acid B. Caro acid C. None of above D. Any of above
9	The sphere is enclosed in brackets in formulas for complex species, and it includes the central metal ion plus the coordinated group	A. Ligand B. Donor C. Coordination D. Oxiation
10	If steel is heated to a temperature well below red heated and is then cooled slowly the process is called.	A. Annealing B. Quenching C. Tempering D. Nitriding
11	In which of the following compounds does hydrogen bonding occur.	A. CCI4 B. NaH C. HI D. NH3
12	The electrolysis of molten metal hydride will produce dihydrogen gas.	A. At cathode B. At anode C. At both the electrodes D. At none of the electrodes
13	Which of the following elements has the highest ionization energy.	A. Na B. Si C. Ar D. Cl
14	The magnitude of electron affinity depends on.	A. Atomic size B. Nuclear charge C. Electronic configuration

	D. All of the above
Type of hybrid orbitals used by the chlroine atom in ClO2 is.	A. sp2 B. sp3 C. sp D. None of these
Which of the following elements has the highest value of IE.	A. Na B. K C. Mg D. Ca
Group III A of the periodic table consist of elemetns.	A. 3 B. 4 C. 5 D. 6
Atomicity of which of the following pair of elements is not same as hydrogen.	A. Phosphorus, Nitrogen B. Nitrogen, Argon C. Nitrogen, iodine D. lodine, sulphur
Titanium dioxide shows the lattice strcuture.	A. Flluorite B. Rutile C. Wurtzite D. Zeolite
The Lewis structure of which of the following does not have coordinate bond.	A. SO2 B. HNO3 C. H2SO4 D. HNO2
	Which of the following elements has the highest value of IE. Group III A of the periodic table consist of elemetrs. Atomicity of which of the following pair of elements is not same as hydrogen. Titanium dioxide shows the lattice struture.