

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Phosphoric acid is the most important of the phosphorus oxy acids. Industrially phosphoric acid is prepared by.	A. The Ostwald process B. The Haber's process C. The reaction of phosphate rock with sulphuric acid D. The reaction P4O10 with water.
2	Which among the following is insoluble in water.	A. LiOH B. KOH C. NaOH D. RbOH
3	Which of the following bonds between carbon -carbon is teh strongest.	A. Sigma bond B. Pi bond C. Double bond D. Triple bond
4	Which among the following is a false statement.	A. SiO2 has a structure similar to that of CO2 B. Natural Si exists only in the combined state C. Si can be prepared by reducing SiO2 with Mg D. Si does not exist in graphite like structure, but exists only ni diamond like structure.
5	The total number of bond pairs around sulphur and total number of lone pairs around oxygen atoms in the Lewis structure of sulphate ion are respectively.	A. 4, 12 B. 8,12 C. 12,4 D. 6 ,12
6	lodine is a grey black solid and its vapours are in color	A. Grey B. Black C. Yellow D. Violet
7	NH3 has a not dipole moment while BF3 has zero dipole moment Thsi is because.	A. NH3 is not a planar molecule while BF3 is a planar molecule. B. NH3 is a planar molecule, while BF3 is a planar molecule. C. Fluorine is more electronetative than nitrogen D. Born is more electronegative than nitrogen
8	Which pair of species can undergo chemical reaction with each other.	A. CO+ NO B. LiH and H2O C. CO2 and HCI D. CaH2 and SI H4
9	Pick out incorrect statement about K2Cr2O7	A. It oxidizes acidified solution H2SO4 to S B. It oxidizes KI to I2 C. It oxidizes HCl to CI2 D. It gives oxygen, when treated with cold conc. H2SO4
10	The concentration of OH- ions in a certain household ammonia solution is 0.0025. This ammonia solution is.	A. Basic B. Acidic C. Neutral D. None of above
11	The pH of the 1.3 x 10 ⁻⁴ NH ₄ Cl is	A. 1.3 B. 4.0 C. 2.886 D. 3.886
12	Which of the following species has highest bond energy.	A. H2 B. T2 C. D D. Cl
		A. Mn is a transition element while K+ is not B. [MnO4]is negatively charged while

13	The atomic number of Potassium is 19 and that of manganese is 25. Although the coloured of MnO4 is dark violet yet the K+ is colourles This is due to the fact that	K+ has a positive charge C. The effective atomicnumebr of Mn is [MnO4] is 26 while for K+ the atomic number is 18 D. The Mn in a high positive oxidation state allows charge transfer transitions
14	Among the elements of third period, the element with lowest boiling point belongs to group.	A. 1 B. 14 C. 16 D. 18
15	Which of the following is not correct.	A. Rusting of iron can be stopped by increasing the concentration of CO2 in water B. Rusting of iron is electrochemical in nature. C. Rusting of iron takes place in moist air D. Rusting of iron produces hydrated iron (III) oxide
16	Which of the following ions is smallest in size.	A. F- B. Cl- C. I- D. Br-
17	In B2H6 molecule	A. There exists a direct B-B a -bond B. All the atoms are in one plane C. All the B-H bonds are normal covalent bonds D. There exist two bonds between the boron atoms.
18	If steel is heated to a temperature well below red heated and is then cooled slowly the process is called.	A. Annealing B. Quenching C. Tempering D. Nitriding
19	Used in filling luminous tubes.	A. Xenon B. Krypton C. Radon D. Helium
20	Which of the following statement is not true regarding Open Hearth process.	A. No iron is lost B. The process is economical and simple C. Steel obtained is of high quality D. Scrap iron cannot be used in this process.