

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	According to SHAH concept the Lewis bases were classified on the basis os.	A. Charge ion size B. Polarization consideration C. Electron and co coordinating ability D. All of above
2	According to SHAB, Lewis acid are divided into.	A. Two classes B. Three classes C. Four classes D. None of above
3	According to Usanovich concept a base is defined as any species.	A. Capable of giving up anions B. Combining with cations C. Neutralizing an acid to give a salt D. All of above
4	Lux -Flood concept is a dono-acceptor system of.	A. Proton B. Electron pair C. Neurtron D. Oxide ion
5	Lewis concept explain the formation of	A. Ionic bond B. Covalent bond C. Co-ordinate bond D. Chemical bond
6	Bases and reducing agents are electron giving agents and also called as.	A. Electrodotic B. Electrophile C. Nucleophile D. None of above
7	The concept is also known as electron pair donor acceptor system.	A. Bronsted Lowery B. Lewis C. Lux -Flood D. Usanovich
8	The concept is also known as proton donor acceptor system.	A. Bronsted Lowery B. Lewis C. Lux Flood D. Usanovich
9	Arrhenius concept explained	A. Constant heat of neutralization B. Quantitative determination of acid base strength C. Catalytic property of acid D. All above
10	"Acids are substance whose aqueous solutions turned blue litmus red and tasted sour" stated by	A. Davy B. Liebig C. Boyle D. Rouelle
11	How many sigma and pi bonds are there in a CO ₂ molecule.	A. 2 sigma B. 2 sigma and 4 pi C. 2 sigma and 2 pi D. 4 sigma and no pi
12	Which of the following interaction is the strong.	A. Dipole -dipole B. Ion induced dipole C. Ion -dipole D. Dipole induced dipole
13	Covalent compound are soluble in	A. Polar solvents B. Non polar solvent C. Concentrated acids D. All solvent
14	The maximum covalence of an element equal to.	A. The number of unpaired d electrons B. The number of paired p electrons C. The number of unpaired a and P electors D. The actual number of a and P electrons in the outermost shell

15	Arrange the following in order of increasing boiling point.	A. CH_3OH < CH_3Cl < RbCl < CH_4 B. CH_3OH < CH_4 < CH_3Cl < RbCl C. RbCl < CH_3Cl < CH_3OH < CH_4 D. CH_4 < CH_3Cl < CH_3OH < RbCl
16	Which of the following statements is wrong.	A. Covalent compounds are generally soluble in polar solvents. B. Covalent compounds have low melting and boiling points. C. Ionic solids do not conduct electricity in solid state. D. Ionic compounds conduct electricity in the fused state.
17	Which of the following statements is wrong.	A. Covalent compounds are generally soluble in polar solvents. B. Covalent compounds have low melting and boiling points. C. Lower than that of separate H atoms. D. Sometimes lower and sometimes higher than that of separate H.
18	Which two atoms of hydrogen combine to form a molecule of hydrogen gas. the energy of the hydrogen molecule is.	A. Higher than that of separate H atoms. B. Equal to that of separate H atoms. C. Lower than that of separate H atoms. D. Sometimes lower and sometimes higher than that of separate H.
19	The boiling point of water is unexpectedly high because.	A. H_2O molecule is linear. B. sp^3 hybridization is involved in the formation of water. C. There is hydrogen bonding and consequent association of H_2O molecules. D. Oxygen is the first member of the VI group.
20	Which of the following statements regarding covalent bond is false.	A. The electrons are shared between atoms. B. The bond is non-directional. C. The strength of the bond depends upon the extent of overlapping. D. The bond formed may be polar or non-polar.