

## PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following bonds will be non polar.	A. N - H B. O - H C. C - H D. Cl - Cl
2	In each period, the most electropositive element belongs to group.	A. 18 B. 17 C. 1 D. 2
3	Ferrochrom contains Cr up to	A. 60-70% B. 70-80% C. 80-90% D. 40-50%
4	An explosive	A. Nitroglycerine B. Trinitrotoluene C. Fluorine perchlorate D. All above
5	Which of the following has the greatest metallic character.	A. Na B. Mg C. Al D. Si
6	The wire of flash bulb is made up of.	A. Cu B. Ag C. Mg D. Ba
7	Variable oxidation states is shown by	A. Normal elements B. Metallic elements C. Non metallic elements D. Transition elements
8	Which of the following would decompose at lowest temperature.	A. MgCO <sub>3</sub> B. SrCO <sub>3</sub> C. BaCO <sub>3</sub> D. CaCO <sub>3</sub>
9	The force responsible for dissolution of ionic compounds in water are	A. Hydrogen bonds B. Ion dipole forces C. Ionic bonds D. Van Der Waal forces
10	The coordination number of closely packed hexagonal is.	A. 4 B. 6 C. 8 D. 12
11	HClO <sub>4</sub> , HNO <sub>3</sub> and HCl are all strong acids in aqueous solution in glacial acetic acid medium, their acid strength is such that.	A. HClO <sub>4</sub> > HCl > HNO <sub>3</sub> B. HNO <sub>3</sub> > HClO <sub>4</sub> > HCl C. HCl > HClO <sub>4</sub> > HNO <sub>3</sub> D. HCl > HClO <sub>4</sub> > HNO <sub>3</sub>
12	The formula of Borax is.	A. Na <sub>2</sub> B <sub>4</sub> O <sub>7</sub> · 6H <sub>2</sub> O B. Na <sub>2</sub> B <sub>4</sub> O <sub>7</sub> · 8H <sub>2</sub> O C. Na <sub>2</sub> B <sub>4</sub> O <sub>7</sub> · 10H <sub>2</sub> O D. Na <sub>2</sub> B <sub>4</sub> O <sub>7</sub> · 12H <sub>2</sub> O
13	Strongest inter molecular hydrogen bond is formed in	A. H <sub>2</sub> O B. NH <sub>3</sub> C. HF D. H <sub>2</sub> S
14	In diborane (B <sub>2</sub> H <sub>6</sub> )	A. The structure is similar to that of C <sub>2</sub> H <sub>6</sub> B. All the atoms are in one plane C. The boron atoms are linked through hydrogen bridges D. There is a direct boron boron bond
		A. SO <sub>2</sub> B. CO <sub>2</sub>

15	Which of the following compounds is electrovalent in nature.	B. ICl C. KBr D. CH <sub>3</sub> I
16	Group III A of the periodic table consist of elemetns.	A. 3 B. 4 C. 5 D. 6
17	The solution of NaOH pH -10.46 contain [OH <sup>-</sup> ]	A. 2.0 X 10 <sup>-4</sup> B. 4.6 X 10 <sup>-4</sup> C. 4.6 X 10 <sup>-2</sup> D. 4.6 X 10 <sup>-3</sup>
18	CCl <sub>4</sub> has zero dipole moment because of.	A. Planar structure B. Tetrahedral structure C. Similar size of C and Cl atoms D. Similar electrons affinity of C and Cl
19	Which of the following bonds between carbon -carbon is teh strongest.	A. Sigma bond B. Pi bond C. Double bond D. Triple bond
20	Which of the following compounds has highest boiling point.	A. HI B. HF C. HBr D. HCl