PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	When 0.01 moles of NaOH are added to a buffer solution, its pH changes from 4.745 to 4.832 WHAT IS ITS.	A. 0.115 B. 0.900 C. 0.015 D. 0.215
2	The pH Value 4.2 is of	A. Vinegar B. Lemons C. Oranges D. Tomatoes
3	Glass electrode cannot be used to measure the pH of pure.	A. Acetic acid B. Ethyl alcohol C. Gelatin D. All above
4	A mixture of weak acid and its salt is.	A. Alkaline buffer B. Acidic buffer C. Neutral buffer D. All of above
5	Buffer solution are used to.	A. Increase the pH B. Resist the pH C. Decrease the pH D. None of above
6	The most convenient and has nearest approach to a universal pH measurement	A. pH strips B. pH indicator C. The emf method D. The colorimetric
7	The solution of NaOH pH -10.46 contain [OH-]	A. 2.0 X. 10⁻⁴ B. 4.6 X 10⁻⁴ C. 4.6 X 10⁻² D. 4.6 X 10⁻³
8	The pH of the 0.0032 M H2SO4 is.	A. 3.2 B. 4.0 C. 2.198 D. 1.0
9	The pH of the 1.3 x 10^{-4} NH_4Cl is	A. 1.3 B. 4.0 C. 2.886 D. 3.886
10	The pH of milk is	A. 6.0 B. 6.5 C. 7.0 D. 7.5
11	The pH of the tears is	A. 7.0 B. 7.4 C. 7.8 D. 8.2
12	The concentration of OH- ions in a certain household ammonia solution is 0.0025. This ammonia solution is.	A. Basic B. Acidic C. Neutral D. None of above
13	The value of Kw increase with temperature because the ionization of water.	A. Positive B. Negative C. Endothermic D. Exothermic
14	The ionic product equilibrium constant is.	A. Ka B. Kb C. Kc D. Kw
15	pH of pure water at 25 ^{o}C. kw = 1 x 10^{-4}	A. 0 B. 7 C. 14 D. None of above

16	A chemical reaction resulting in a change in the electric charge on the reacting particles may be called as.	A. Add ion reaction B. Redox reaction C. Elimination reaction D. Chain reaction
17	Relative order of acidity of oxy acid	A. HClO > HClO2 > HClO3 > HClO4 B. HClO4 > HClO3 >HClO2 > HClO C. HClO3 > HClO2 > HClO > HClO4 D. HClO2 > HClO4 > HClO3 > HClO
18	Relative order of acidity of HF, HCl, HBr, and HI acids is	A. HCl > HBr > HI > HF B. HF > HCl > HBr > HI C. HI > HBr > HCl >HF D. HF > HI > HCl > HBr
19	All the strong acids have very close pKas value and they appear to have nearly equal strengths in aqueous solutions. The phenomenon is called as.	A. Levelling effect B. Differnetiating effect C. Levelling solvent D. Differnetiating solvent
20	The one in which the acceptor atom is of low positive charge, Large size and has several outer electrons which can be easily excited is a.	A. Soft base B. Hard Base C. Soft acid D. Hard acid