PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	After assimilation urea leaves behind in the soil	A. NH3 B. CO2 C. Both A and B D. None of above
2	Ca2+ is isolelectronic with.	A. Mg2+ B. Kr C. Ar D. Na+
3	Which of the following elements display maximum tnedency to form P Pi - p PI multiple bonds with itself and with carbon and oxygen.	A. N B. p C. Bi D. As
4	The _____ sphere is enclosed in brackets in formulas for complex species, and it includes the central metal ion plus the coordinated group	A. Ligand B. Donor C. Coordination D. Oxiation
5	The ion that is isoelectornic with CO is	A. CN- B. O2+ C. CO2- D. N2+
6	Which of the following are neutral ligands.	A. NH3 B. H2O C. CO & NO D. All of above
7	The formula of Tetraboric acid is.	A. H2BO3 B. HBO2 C. H2B4O7 D. H6B4O9
8	Which of the following is a pseudohalide.	A. I3- B. IF7 C. CN- D. ICl
9	The electronic configuration of some elements are given below. The element with highest electron affinity is	A. 1s2, 2s2, 2p3 B. 1s2, 2s2, 2p4 C. 1s2, 2s2, 2p5 D. 1s2, 2s2, 2p6
10	The atomic number of potassium is 19 and that of mangness is 25 Although the colour of MnO4 is dark violet yet the K+ is colourless this is due to the fact that.	A. Mn is a transition element while K+ is not B. [MnO4]- is negatively charged while K+ has positive charge C. The effective atomic number of Mn is [MnO4]- is 26: while for K+ the atomic number is 18 D. The Mn is a high positive oxidation state allows charge transfer transitions.
11	Fluorine does not show variable oxidation state because of.	A. its high electronegativity B. Its small sixe C. low dissociation energy of F-F bond D. Non availability of d-orbitals
12	Sodium silicate is ued	A. In fire proofing of wood and textiles B. As a preservative of eggs C. As a furniture polish D. All above
13	In extraction of iron, the furnace charge consists of iron ore, coke and limestone. The function of limestone is to act as.	A. An oxidation agent B. A reducing agent C. Flux D. Slag
14	The most abundant metal in earth's crust is.	A. Fe B. Al C. Ti

		D. Ca
15	For covalent bond to form between two atoms A and B	A. Transference of electrons must take place from A to B B. A pair of electrons of A is shared by both A and B C. A and B contribute equal no. of electrons for mutual sharing by A and B D. One of the atom A or B must already have octet of electrons.
16	The increase in boiling points of noble gases from He to Xe is due to the	A. Decreases in ionization energy B. Increases in polarizability C. Increase in electron affinity D. Increase in atomic volume
17	The noble gases which does not I do not form any clathrates is.	A. He B. Ne C. Argpm D. Both He and Ne
18	The interactions in HF are.	A. dipole dipole interasctions B. Hydrogen bonds C. dipole -dipole and dispersion forces D. Hydrogen bond and diapersin forces
19	Which element out of the following can ehibit a maximum con valency of seven.	A. Chlorine B. Sulphur C. Fluorine D. both Cl and F
20	Which of the following statements is wrong.	A. Covenant compounds are generally soluble is polar solvents. B. Covalent compounds have low melting and boiling points C. Lower than that of separate H atoms D. Sometimes lower and sometimes higher than that of separate H