

PPSC Chemistry Part III Inorganic Chemistry Online Test

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Sr	Questions	Answers Choice
1	The atomic number of potassium is 19 and that of mangness is 25 Although the colour of MnO4 is dark violet yet the K+ is colourless this is due to the fact that.	A. Mn is a transition element while K+ is not B. [MnO4]- is negatively charged while K+ has positive charge C. The effective atomic number of Mn is [MnO4]- is 26: while for K+ the atomic number is 18 D. The Mn is a high positive oxidation state allows charge transfer transitions.
2	Anhydrous AICl3 cannot be obtained by heating hydrated AIOl3 ,6h2o Because.	A. It decomposes completely to give Al2O3 B. It does not lose water completely C. It undergoes hydrolysis to give Al(OH)3 D. AlCl3 .6H2O is very stable.
3	The most stable oxidation state of chromium is.	A. +6 B. +3 C. +4 D. +2
4	Which of the following is not an ore of iron.	A. Haematite B. Magnetite C. Siderite D. Monazite
5	Sulphate ores of aluminium	A. Alumite B. Cryolite C. Fekdsper D. Kaolin
6	The noble gases are found in the atmosphere to the extent of about some percent by volume.	A. 0.5% B. 1.0% C. 1.5% D. 2.0%
7	Pick out incorrect statement about K2Cr2O7	A. It oxidizes acidified solution H2SO4 to S B. It oxidizes KI to I2 C. It oxidizes HCI to CI2 D. It gives oxygen, when treated with cold conc. H2SO4
8	A boy accidently splashes a few drops of conc. H2SO4 on his cotton shirt. A few minutes later, the splashed part blacken and holes appear. This is because the sulphuric acid.	A. Heats up the cotton so that it burns B. Dehydrates the cotton C. Causes cotton to react with oxygen of the air D. Removes the elements of water from cotton
9	By applying an external force the ionic solid can be easily broken to powder form so the ionic solid are highly	A. Hard B. Brittle C. Tough D. Soft
10	The interactions in HF are.	A. dipole dipole interasctions B. Hydrogen bonds C. dipole -dipole and dispersion forces D. Hydrogen bond and diapersin forces
11	Relative order of acidity of oxy acid	A. HCIO > HCIO2 > HCIO3 > HCIO4 B. HCIO4 > HCIO3 >HCIO2 > HCIO C. HCIO3 > HCIO2 > HCIO > HCIO4 D. HCIO2 > HCIO4 > HCIO3 > HCIO
		A. There exists a direct B-B a -bond

12	In B2H6 molecule	B. All the atoms are in one plane C. All the B-H bonds are normal covalent bonds D. There exist two bonds between the boron atoms.
13	Which of the following pairs shows diagonal relationship	A. Li and Mg B. Na and K C. Zn and Cd D. Li and BE
14	Lewis concept explain the formation of	A. lonic bond B. Covalent bond C. Co-ordinate bond D. Chemical bond
15	Which of the following statement is not true for carbon.	A. Its forms compounds with multiple bonds B. Its ionization energy is very high C. It undergoes catanation D. It shows inert pair effect
16	The configuration of valence shell of certain atom X is $3s2$, $3p5$, which valences can it exhibit.	A. 1,3 only B. 1,5 only C. 1,3,5,7 D. 1,3,4
17	Halogens are coloured because.	A. They are strong oxidant B. Their molecules are held together by weak van der Waals forces C. Their atoms absorb radiations form visible range causing the excitation of valence electrons to higher energy of levels D. Their molecules absorb light radiation forming the oxcited state.
18	lonic reactions mainly take place in.	A. Aqueous solutions and organic solvents of high polarity B. Non aqueous solvents of low polarity C. Gaseous state D. Solid state
19	Major achievement of CFT is	A. Interpreting the color B. Adsorption spectra C. Both A and B D. None of above
20	Boric acid is added to glass because is.	A. Makes the glass opalescent B. Reduces the coefficient of expansion C. Makes the glass brittle D. Increase refractive index of the glass.