

## PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following statements is correct.	A. A sigma bond is weaker than a pi bond B. There are four coordinate bonds in the Lewis structure of $\text{NH}_4^+$ ion. C. The 1 covalent bond is directional in nature D. A single bond between the two atoms cannot be re bond.
2	Which element out of the following can exhibit a maximum co valency of seven.	A. Chlorine B. Fluorine C. Sulphur D. Both Cl and F
3	Reacts violently with water	A. $\text{AlH}_3$ B. $\text{AlCl}_3$ C. $\text{LiAlH}_4$ D. $\text{Al}_2\text{Cl}_6$
4	The bond angle along $\text{Sp}^2$ hybridization is.	A. $180^\circ$ B. $120^\circ$ C. $109.5^\circ$ D. $160^\circ$
5	Setting of cement is improved by	A. Lime stone B. Clay C. Gypsum D. Water
6	The structure of $\text{SO}_2$ is	A. Linear B. Angular C. V-shaped D. Planner
7	According to the VSEPR theory, the shape of the $\text{SO}_3$ molecule is.	A. Pyramidal B. Tetrahedral C. Trigonal planar D. Distorted tetrahedron
8	Bases and reducing agents are electron giving agents and also called as.	A. Electrodotic B. Electrophile C. Nucleophile D. None of above
9	Which is not an ore of aluminium.	A. Bauxite B. Cryolite C. Monazite D. Corundum
10	A trend which is common to elements of both the group IA and group VII A ongoing from top to bottom.	A. Boiling point increases B. Electron affinity increases C. Oxidizing power increases D. Ionization energy decrease
11	The halide which is inert to water is	A. $\text{PCl}_5$ B. $\text{SiCl}_4$ C. $\text{BCl}_3$ D. $\text{SF}_3$
12	The element having highest ionization energy and least electron affinity belong to	A. Period 1 , group 18 B. Period 2, group 17 C. Period 2, group 1 D. Period 2, group 2
13	Which of the following elements would have the lowest first ionization energy	A. Mg B. Rb C. Li D. Ca
14	The state of hybridization of carbon in $\text{CO}_2$ is	A. $\text{sp}^2$ B. $\text{sp}$ C. $\text{sp}^3$ D. $\text{dsp}^2$

15 The rusting of iron is catalyzed by which of the following.  
A.  $\text{He}$   
B.  $\text{H}^+$   
C.  $\text{O}_2$   
D.  $\text{Zn}$

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16 Which of the following elements display maximum tendency to form  $\text{P}_\text{Pi} - \text{p}_\text{Pi}$  multiple bonds with itself and with carbon and oxygen.  
A. N  
B. p  
C. Bi  
D. As

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17 Which pair of species can undergo chemical reaction with each other.  
A.  $\text{CO} + \text{NO}$   
B.  $\text{LiH}$  and  $\text{H}_2\text{O}$   
C.  $\text{CO}_2$  and  $\text{HCl}$   
D.  $\text{CaH}_2$  and  $\text{SiH}_4$

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18 VBT does not explain  
A. Absorption spectra  
B. Color of transition metal ion  
C. Heat of formation  
D. All above

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19 Which of the ionic possesses highest bond energy.  
A. C-C  
B. Si-Si  
C. Ge-Ge  
D. Sn-Sn

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20 The valence shell electronic configuration of group III A is.  
A.  $\text{ns}1 \text{ p}2$   
B.  $\text{ns}2 \text{ p}1$   
C.  $\text{ns}3\text{p}2$   
D.  $\text{ns}2\text{p}2$