

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following alkaline earth metals occurs in radioactive form in nature.	A. Ca B. Mg C. Ra D. Ba
2	Which of the following metals is the most abundant in the earth's crust.	A. Mg B. Ca C. K D. Na
3	Which of the following hydroxide is amphoteric.	A. $\text{B}(\text{OH})_3$ B. $\text{Al}(\text{OH})_3$ C. $\text{Ga}(\text{OH})_3$ D. $\text{In}(\text{OH})_3$
4	When 0.01 moles of NaOH are added to a buffer solution, its pH changes from 4.745 to 4.832. WHAT IS ITS.	A. 0.115 B. 0.900 C. 0.015 D. 0.215
5	Which of the following process is used for the conversion of matte is to nickel.	A. Orford process B. Mond's process C. Electrolytic process D. All
6	Which of the following is not a property of Ni.	A. it is a soft silvery white metal B. It is malleable and ductile C. It is highly magnetic D. It has high electrical and thermal conductivities
7	The strongest acid is.	A. HNO_2 B. HNO_3 C. $\text{H}_2\text{N}_2\text{O}_2$ D. HNO_5
8	In which of the following characteristics does hydrogen resemble halogens.	A. Hydrogen is the lightest gas B. Hydrogen contains one electron each C. Hydrogen forms ionic hydrides with alkali metals D. Hydrogen has three isotopes.
9	Which of the following halide has lowest melting point.	A. NaCl B. NaF C. NaBr D. NaI
10	Among the elements of second period the element with highest melting point belongs to group.	A. 1 B. 14 C. 17 D. 18
11	Molecules have zero dipole moment	A. CO_2 B. BCl_3 C. CH_4 & CCl_4 D. All above
12	The name hydrogen was proposed by.	A. Lavoisier B. Rutherford C. Henry Cavendish D. Scheele
13	Which among the following hydride is ionic in nature.	A. Ammonia B. Protium oxide C. Calcium hydride D. Sulphane
14	Aluminium is used for.	A. Making utensiles & frame B. Making alloys C. Reducing agent D. All above
		A. One - -

15	Excluding H-atom , Hydrogen bond never involves more than atoms.	B. Two C. Three D. Four
16	Berllium has diagonal relationship with	A. Li B. Al C. B D. Na
17	Of the following the commonly used in the laboratory desiccator is.	A. Anhyd. Na ₂ Co ₃ B. Anhyd. Ca Cl ₂ C. Dry NaCl D. None of the above
18	The oxidation state shown by phosphorus is.	A. - 3 B. + 3 C. + 3 and +5 D. -3 ,+ 3 and +5
19	In which pair of species, the Lewis formula contain same number of Lone pairs and bond pairs but they are not iso electronci.	A. O ₂ B ₂ B. SO ₂ , O ₃ C. PCI ₃ , BF ₃ D. SOCl ₂ , COCl ₂
20	Electron gas theory fails to explain	A. Specific heat of metals B. Electrical and thermal conductivity C. Paramagnetic behavior of metals D. All of the above