

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Stable metal ions structures are.	A. Noble gas structure B. Is electron group structure C. Transition metal in structure D. All of the above
2	The type of bonding in HCl is	A. Pure covalent B. Polar covalent C. Highly polar D. Hydrogen bonding
3	Strongest inter molecular hydrogen bond is formed in	A. H ₂ O B. NH ₃ C. HF D. H ₂ S
4	Mangalium is an alloy of.	A. Al + Mg B. Mg + Al + Mn C. Mg + Al + Cu D. Mg + Al + Cu + Mn
5	Glass electrode cannot be used to measure the pH of pure.	A. Acetic acid B. Ethyl alcohol C. Gelatin D. All above
6	When orthoboric acid is heated strongly it gives.	A. B ₂ O ₃ B. H ₂ B ₃ O ₇ C. HBO ₂ D. B
7	Which of the following is an allotrops of hydrogen.	A. 0- H ₂ B. P-H ₂ C. Both A and B D. None of these
8	The noble gases are found in the atmosphere to the extent of about some percent by volume.	A. 0.5% B. 1.0% C. 1.5% D. 2.0%
9	The Lewis formula of SOCl ₂ , the total number of bond pairs and lone pairs of electron around sulphur are.	A. 2,1 B. 2,2 C. 3,1 D. 3,0
10	Aluminium halides is.	A. White crystalline solid B. Hygroscopic C. Sublimes at 180°C D. All above
11	The pH of the 0.0032 M H ₂ SO ₄ is.	A. 3.2 B. 4.0 C. 2.198 D. 1.0
12	Helium is used in weather balloons and airships instead of H ₂ because it is.	A. Lighter than hydrogen B. Incombustible C. More abundant than hydrogen D. Radiative
13	The basic strength of hydrides of group 15 elements vary in the following order.	A. NH ₃ > PH ₃ > AsH ₃ > SbH ₃ > BiH ₃ B. PH ₃ > NH ₃ > AsH ₃ > SbH ₃ > BiH ₃ C. BiH ₃ > NH ₃ > PH ₃ > AsH ₃ > SbH ₃ D. NH ₃ > PH ₃ > SbH ₃ > AsH ₃ > BiH ₃
14	Which of the following is diamagnetic	A. O ₂ B. O ₂ ⁺ C. O ₂ ⁻ D. O ₂ ²⁻

15	Which of the following is not a property of Ni.	<p>A. It is a soft silvery white metal</p> <p>B. It is malleable and ductile</p> <p>C. It is highly magnetic</p> <p>D. It has high electrical and thermal conductivities</p>
16	Ground state electronic configuration of valence shell in N_2 molecule is written as $(\sigma 2s)^2, (\sigma^* 2s)^2, (\pi sp)^4, (\sigma 2p)^2$. Hence, the bond order of N_2 molecule is.	<p>A. 1</p> <p>B. 2</p> <p>C. 3</p> <p>D. 0</p>
17	The one in which the acceptor atom is of low positive charge, Large size and has several outer electrons which can be easily excited is a.	<p>A. Soft base</p> <p>B. Hard Base</p> <p>C. Soft acid</p> <p>D. Hard acid</p>
18	Chlorine is used in	<p>A. Sterilization of water</p> <p>B. Extraction of gold</p> <p>C. Bleaching of cotton</p> <p>D. All above</p>
19	The element having highest ionization energy and least electron affinity belong to	<p>A. Period 1, group 18</p> <p>B. Period 2, group 17</p> <p>C. Period 2, group 1</p> <p>D. Period 2, group 2</p>
20	Which of the following element has six electrons in the valence shell but cannot exhibit a maximum covalency of six.	<p>A. Sulphur</p> <p>B. Oxygen</p> <p>C. Selenium</p> <p>D. Both A and B</p>