

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	The common oxidation state of elements of group V A is.	A. -3 B. +3 C. +5 D. Any above
2	Noble gases are used in discharge tubes to give different colours. Raddish orange glow is due to.	A. Ar B. Ne C. Xe D. Kr
3	Which among the following is insoluble in water.	A. LiOH B. KOH C. NaOH D. RbOH
4	"Acids are substance whose aqueous solutions turned blue litmus red and tasted sour" stated by	A. Davy B. Liebig C. Boyle D. Rouelle
5	Which is not an ore of aluminium.	A. Bauxite B. Cryolite C. Monazite D. Corundum
6	The yellow colour of chromates changes to orange red on acidification, due to the formation of.	A. Cr ³⁺ B. Cr ₂ O ₃ C. Cr ₂ O ₇ ²⁻ D. Cro ₃
7	The acetylene molecule contain a	A. Single bond B. Double bond C. Triple bond D. Co ordinate bond
8	The penultimate shell of carbon contains electrons.	A. s ₂ B. s ₂ p ₆ C. s ₂ p ₆ d ₁₀ D. s ₂ p ₆ d ₈
9	The geometry of the molecule is primarily decided by	A. Bond pairs around the central atom B. No of k bond around the central atom C. No of bond pairs as well as lone pairs around the central atom D. No. of lone pairs on central atom
10	Copper occurs in nature as.	A. Native B. Combined C. Both native and combined D. None of the above
11	Which among the following is a false statement.	A. SiO ₂ has a structure similar to that of CO ₂ B. Natural Si exists only in the combined state C. Si can be prepared by reducing SiO ₂ with Mg D. Si does not exist in graphite like structure, but exists only ni diamond like structure.
12	The hydrogen bond is strongest in.	A. O - HS B. S - H.....O C. F- H.....F D. F - H.....O
13	Which element are non metals.	A. N & P B. Sb & Bi C. As & Sb D. N & Bi
14	H-Bonding also exists in living system like	A. Protein B. DNA C. Carbohydrates D. All of these

	Increasing and decreasing systems	<p>C. Both A and B</p> <p>D. None of above</p>
15	In the electronic structure of acetic acid, the total number of shared and unshared pair of electrons are respectively.	<p>A. 16, 8</p> <p>B. 8, 4</p> <p>C. 12, 8</p> <p>D. 8, 12</p>
16	The first ionization energies of the elements of the first transition series. (Ti _____ Cu)	<p>A. Increases as the atomic number increases</p> <p>B. decreases as the atomic number increases</p> <p>C. Do not show any change as the addition of electrons takes place in the inner (n-1) d-orbitals.</p> <p>D. Increases from Ti to Mn and then decreases from Mn to Cu</p>
17	Which of the following hydroxides has the maximum solubility in water.	<p>A. Mg (OH)₂</p> <p>B. Ca (OH)₂</p> <p>C. Sr (OH)₂</p> <p>D. Ba (OH)₂</p>
18	Which element out of the following can exhibit a maximum co valency of seven.	<p>A. Chlorine</p> <p>B. Fluorine</p> <p>C. Sulphur</p> <p>D. Both Cl and F</p>
19	Bromine is used as	<p>A. Fungicides</p> <p>B. Herbicides</p> <p>C. Germicides</p> <p>D. Insecticides</p>
20	The types of coordinate compounds.	<p>A. Labile</p> <p>B. Inert</p> <p>C. Both A and B</p> <p>D. None of above</p>