

## PPSC Chemistry Part III Inorganic Chemistry Online Test

0	Overtions	Arrayana Oh. :
Sr	Questions	Answers Choice
1	Carbon in wrought iron is present as	A. Silicon carbide     B. Iron carbide or comentite     C. Graphite     D. Partly as iron carbide and partly as graphite
2	According to Usanovich concept a base is defined as any species.	<ul><li>A. Capable of giving up anions</li><li>B. Combining with cations</li><li>C. Neutralizing an acid to give a salt</li><li>D. All of above</li></ul>
3	Ground state electronic configuration of valence shell in N2 m lecule is written as (a2s)2, (o*2s)2, (pi sp)4, (o2p)2, Hence, the bond order of N2 molecule is.	A. 1 B. 2 C. 3 D. 0
4	The first ionization energy in electron volts of nitrogen and oxygen atoms are respectively given by.	A. 14.6, 13.6 B. 13.6, 14.6 C. 13.6, 13.6 D. 14.6, 14.6
5	Which of the following statements is not correct about noble gases.	A. Their ionization energies are very high B. Their electron affinities are nearly zero C. They do not form any chemical compounds D. They are not easily liquefied
6	Phosphorus normally exhibit a covalency of.	A. +1 and +2 B. +2 and +3 C. +3 and +4 D. +4 and +5
7	Which of the following statement is not correct.	A. The element with highest IE belongs to group 18 B. In each period the element with lowest IE belongs to group I C. In each period the element with highest IE is a noble gas D. In the second period as we move from left to right, ionization energy increases regularly.
8	The common oxidation state of lanthanides is.	A. +3 B. +2 C. +1 D. +4
9	The state of hybridization of Xe in Xe F6 are	A. sp2 B. sp3 C. sp3 d D. dsp3
10	In 1952 who popularized the use of CFT for inorganic chemist	A. Bethe B. Orge C. Van Vleck D. Werner
11	Arrange the following in order of increasing boiling point.	A. CH3OH ⁢ CH3CI ⁢ RbCI ⁢ CH4 B. CHOH ⁢ CH4 ⁢ CH3CI ⁢ RbCI C. RbCI ⁢ CH3CI ⁢ CH3OH ⁢ CH4 D. CH4 ⁢ CH3CI ⁢ CH3OH ⁢ RbCI
12	Metal crystallize is system having co ordination number	A. 8 B. 12 C. 14 D. any one of above
13	The three isotopes of hydrogen differ from one another in	A. Atomic number B. Number of protons

	C. Nuclear charge D. Nuclear mass
Copper occurs in nature as.	A. Native B. Combined C. Both native and combined D. None of the above
A covalent bond which is formed between two atoms by the overlap of atomic orbitals along their axis is called.	A. Pi bond B. Sigma bond C. Polar bond D. Non polar bond
In the fourth flotation process for the purification of ores, the ore particles floats because.	A. They are light B. Their surface is not easily wetted by water C. They bear electrostatic charge D. They are insoluble
In the Lewis structure of H2SO4 molecule the total number of unshared electrons in valence shell of various atoms is.	A. 8 B. 16 C. 12 D. 20
Greenish yellow gas with pungent irritating odour	A. Chlorine B. Fluorine C. lodine D. Bromine
ls a peroxy acid	A. H2SO5 B. H2S2O6 C. H2SO4 D. H2S2O7
How many sigma and pi bonds are there in a CO2 molecule.	A. 2 sigma B. 2 sigma and 4 pi C. 2 sigma and 2 pi D. 4 sigma and no pi
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