PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	The commonly used catalyst in the manufacture of H2SO4	A. Fe2O3 with a little CuO B. v2O5 C. Platinized asbestos and MgSO4 D. All above
2	Molecule of oxygen is	A. Diamagnetic B. Paramagnetic C. Both A and B D. None of above
3	Green houses are responsible for keeping our plant warm and sustaining life on the earth.	A. CO2 & water vapours B. CO2 & CFC C. CO2 & H2O D. CO2 & CH4
4	Metallic magnesium is obtained by	A. Reduction of MgO with Coke B. Electrolysis of an aqueous solution of MgCl2 C. Electrolysis of molten MgCl2 D. Displacement of magnesium by iron form MgCl2 solution.
5	Which of the following statements about anhydrous aluminium chloride is correct.	A. It exist as AlCl3 molecules B. It is not easily hydrolysed C. It sublimes at 100 ^oC under vacuum D. Boron does not form B³⁺ ions
6	Aluminum is usually extracted from	A. Bauxite B. Corundum C. Feldepar D. Alumite
7	A type of a chemical bond which is formed by the mutual sharing of electrons between combining atoms of the same or different elements is called.	A. Ionic bond B. Covalent bond C. Co ordinate Covalent bond D. Metallic bond
8	Pick out the incorrect statement.	A. Red phosphorus consists of a complied chain structure and black phosphorus has a layer structure. B. Nitrogen shows a little tendency for catenation, because N-N a single bond is very strong. C. The xamimum number of covalent bonds formed by nitrogen is four, since it has no d-orbitals in its valence shell D. The group 15 elements do not form M5+ ions, but +5 oxidatin state is realized only through covalent bonding.
9	The number of electrons involved in bonding in Lewis structure of oxalate ion is	A. 20 B. 14 C. 22 D. 18
10	The unequal sharing of bonded pair of electrons between the two atoms in a molecule causes.	A. Dipole B. Radical formation C. Decomposition of found D. Covalent found
11	The atomic number of Potassium is 19 and that of manganese is 25. Although the coloured of MnO4 is dark violet yet the K+ is colourles.. This is due to the fact that	A. Mn is a transition element while K+ is not B. [MnO4]is negatively charged while K+ has a positive charge C. The effective atomicnumebr of Mn is [MnO4] is 26 while for K+ the atomic number is 18 D. The Mn in a high positive oxidation state allows charge transfer transitions

12	Gases and dust particles are removed from H_2SO_4 by	A. Tydal effect B. Drying tower C. Absorption tower D. Contact converter
13	Which compound among the following does not contain an ionic bond.	A. NaOH B. HCl C. KaS D. LiH
14	The common ligands can be arranged in order of their increasing splitting power to cause d-orbitals splitting. This series is called as.	A. Electro-chemical B. Spectro -chemical C. Physico-chemical D. Spectro -electrical
15	Ionization potential of carbon is.	A. 11.2 B. 7.8 C. 8.1 D. 7.3
16	Flourine differs from the other members of its own group due to.	A. Its small size and low bond energy B. Its higher electornegativity C. None-availability of d-orbitals in its valence shell D. All the above
17	Electronegativity of oxygen is.	A. 2,5 B. 3,5 C. 2,4 D. 2.1
18	Which of the following ions is smallest in size.	A. F- B. Cl- C. I- D. Br-
19	Is a chain silicate	A. Olivine B. Tremolite C. Beryl D. Zeolite
20	Sulphate ores of aluminium	A. Alumite B. Cryolite C. Fekdsper D. Kaolin