

## PPSC Chemistry Part III Inorganic Chemistry Online Test

| Sr | Questions   | Answers Choice   |
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| 1  | Nitric acid is used in manufacturing of.  | A. Explosive<br>B. H <sub>2</sub> SO <sub>4</sub><br>C. Fertilizer<br>D. All above   |
| 2  | H <sub>2</sub> SO <sub>4</sub> is manufactured by   | A. The lead chamber process<br>B. The contact process<br>C. Both A and B<br>D. The Ostwald's process   |
| 3  | The noble gases are used due to having property   | A. Chemical inertness<br>B. Low boiling point<br>C. Any of a or b<br>D. Both a and b   |
| 4  | Elements in the same vertical group of the periodical have same                           | A. Number of electron<br>B. Atomic number<br>C. Number of valence electrons<br>D. Electronic configuration   |
| 5  | Molecule of oxygen is   | A. Diamagnetic<br>B. Paramagnetic<br>C. Both A and B<br>D. None of above   |
| 6  | Which among the following is a false statement.   | A. SiO <sub>2</sub> has a structure similar to that of CO <sub>2</sub><br>B. Natural Si exists only in the combined state<br>C. Si can be prepared by reducing SiO <sub>2</sub> with Mg<br>D. Si does not exist in graphite like structure, but exists only in diamond like structure. |
| 7  | CFSE for d <sup>7</sup> ion is.   | A. 0.8<br>B. -0.8<br>C. -1.8<br>D. 1.8   |
| 8  | In urea the amount of nitrogen is   | A. 82.0%<br>B. 46.0%<br>C. 33.0%<br>D. 21.0%   |
| 9  | Any substance which has solidified from the liquid state with crystallization is known as | A. Steel<br>B. Fibre<br>C. Glass<br>D. Asbestos  |
| 10 | Since the acid gives both acidic and normal salts so the acid is.                         | A. di acid<br>B. di basic<br>C. double salt<br>D. Any of above   |
| 11 | Stainless steel contains  | A. Fe+Cr+Ni<br>B. Fe+Ni+Cu<br>C. Fe + Cr+ Cu<br>D. Cu + C + Ni   |
| 12 | Elements of group 14 have the electronic configuration of their outer shell as            | A. ns <sup>2</sup> np <sup>3</sup><br>B. ns <sup>2</sup> np <sup>2</sup><br>C. ns <sup>2</sup> np <sup>6</sup><br>D. ns <sup>2</sup>   |
| 13 | The magnitude of electron affinity depends on.  | A. Atomic size<br>B. Nuclear charge<br>C. Electronic configuration<br>D. All of the above  |
| 14 | Which one of the following has the biggest electron affinity.                             | A. F <sub>2</sub><br>B. Cl <sub>2</sub><br>C. Br <sub>2</sub><br>D. I <sub>2</sub>   |

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| 15 | The penultimate shell of carbon contains electrons.         | A. s <sup>2</sup><br>B. s <sup>2</sup> p <sup>6</sup><br>C. s <sup>2</sup> p <sup>6</sup> d <sup>10</sup><br>D. s <sup>2</sup> p <sup>6</sup> d <sup>8</sup> |
| 16 | Which of the following has the highest melting poing.       | A. NaCl<br>B. Li Cl<br>C. KCl<br>D. Rb Cl  |
| 17 | Nitric acid has the property                                | A. <div>Nitrating</div><br>B. Reducing<br>C. Redoxing<br>D. None of above  |
| 18 | Which of the ionic possesses highest bond energy.           | A. C-C<br>B. Si -Si<br>C. Ge - Ge<br>D. Sn -Sn   |
| 19 | Which of the following compounds has highest boiling point. | A. HI<br>B. HF<br>C. HBr<br>D. HCl   |
| 20 | Sodium silicate is ued                                      | A. In fire proofing of wood and textiles<br>B. As a preservative of eggs<br>C. As a furniture polish<br>D. All above   |