

PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following have identical bond order.	A. CN ⁻ and O ₂ ⁻ B. CN ⁻ and NO ⁺ C. O ₂ ⁻ and CN ⁺ D. NO ⁺ and CN ⁺
2	The electronic configuration of some elements are given below. The element with highest electron affinity is	A. 1s ² , 2s ² , 2p ³ B. 1s ² , 2s ² , 2p ⁴ C. 1s ² , 2s ² , 2p ⁵ D. 1s ² , 2s ² , 2p ⁶
3	The three dimensional silicate anion (SiO ₅ ²⁻) _n is present in	A. Beryl B. Silica C. Asbestos D. Clays
4	VBT does not explain	A. Absorption spectra B. Color of transition metal ion C. Heat of formation D. All above
5	Diborane is used	A. For high energy fuel B. For welding torches C. as reducing agent D. All above
6	Sodium react more vigorously than lithium because.	A. It is a metal B. It has higher atomic mass C. It is more electronegative D. It is more electropositive
7	Ground state electronic configuration of valence shell in N ₂ molecule is written as (σ [*] 2s) ² , (π sp) ⁴ , (σ2p) ² , Hence, the bond order of N ₂ molecule is.	A. 1 B. 2 C. 3 D. 0
8	Bromine is soluble in	A. Alcohol B. Water C. Chloroform D. All above
9	In which pair of species, the Lewis formula contain same number of Lone pairs and bond pairs but they are not iso electronic.	A. O ₂ , B ₂ B. SO ₂ , O ₃ C. PCI ₃ , BF ₃ D. SOCl ₂ , COCl ₂
10	Commercial or the phosphoric acid is pure.	A. 37.0% B. 82.98% C. 88.25% D. 90.12%
11	The central metal atom or ion and the ligands that are directly attached to it are enclosed in a square bracket called.	A. Coordination complex B. Coordination sphere C. Coordination number D. Coordination compounds
12	The following oxo acids have been arranged in the order decreasing acid strength identify the correct order.	A. III > IV > II > I B. III > II > I > IV C. I > II > III > IV D. IV > III > II > I
13	Noble gases are sparingly soluble in water owing to.	A. Dipole -dipole interactions B. Dipole -induced dipole interactions C. Hydrogen bonding D. Induced dipole -instantaneous dipole interactions
14	In each period the element with least electron affinity belongs to.	A. Group 1 B. Group 14 C. Group 17 D. Group 18
15	Which of the following compounds liberates CO ₂ on heating.	A. Li ₂ CO ₃ B. Na ₂ CO ₃ C. K ₂ CO ₃ D. NH ₄ CO ₃

D. All liberate CO₂ on heating.

16 pH of pure water at 25 °C. $K_w = 1 \times 10^{-14}$

- A. 0
- B. 7
- C. 14
- D. None of above

17 Explosive trioxide XeO₃ is produced when

- A. XeOF₄ reacts with water
- B. XeOF₄ reacts with silica
- C. XeF₄ reacts with water
- D. Any of above statements

18 Which one of following is non polar

- A. CH₂Cl₂
- B. CCl₄
- C. CHCl₃
- D. CH₃Cl

19 Select the correct IUPAC name for [Co(NH₃)₆]²⁺

- A. Hexamminiacobaltate (II) ion
- B. Hexaammincobaltate (II) ion
- C. Hexamminiacobalt (II) ion
- D. Hexaammincobalt (II) ion

20 Helium is used for

- A. The preservation of food
- B. Filling electrical transformer
- C. Pressuring agent in rockets
- D. All of above