

## PPSC Chemistry Part II Organic Chemistry Online Test

Sr	Questions	Answers Choice
1	The bond angle between hybrid orbitals in methane is	A. 115.5 <sup>o</sup> B. 109.5 <sup>o</sup> C. 105.7 <sup>o</sup> D. 120 <sup>o</sup>
2	Which of the following does not belong in the group of herocyclic dyes.	A. Acridine B. Cyanine C. Methylene blue D. Amido black
3	Monomer of natural rubber is	A. 1,3-Butadiene B. 2-Methyl -1,3-butadiene C. 1,2 -Butadiene D. 1,3 - Pentadiene
4	The unit of nucleic acid having base sugar combination is called.	A. Nucliec acid B. Nucleoside C. Nucleotide D. None of these
5	The electrophile in the sulphonation of benzene is.	A. SO <sub>3</sub> B. SO <sub>3</sub> H C. HSO <sub>4</sub> D. SO <sub>2</sub>
6	Which of the following is not a general property of amino acids.	A. They have high m.p. and b.p B. They are soluble in water C. Their dipole moments are high D. They are amorphous solids
7	Which of the following hydrocarbon cannot be obtained on reacting chloomethane with sodium metal in the presence of dry ether.	A. C <sub>4</sub> H <sub>10</sub> B. C <sub>2</sub> H <sub>6</sub> C. C <sub>2</sub> H <sub>4</sub> D. C <sub>3</sub> H <sub>8</sub>
8	Which of the following case of acid or base strength is not explained by inductive effect.	A. Formic acid &gt; acetic acid B. Dimethyl amine &gt; trimethyl amine C. Dimethyl amine &gt; methyl amine D. Chloroacetic acid &gt; acetic acid
9	The condensation between formaldehyde and acetaldehyde in the presence of conc. NaOH and heat gives.	A. Acrolein B. Mixture of CH <sub>3</sub> OH and CH <sub>3</sub> COO Na. C. Mixture of CH <sub>3</sub> CH <sub>2</sub> OH and HCOO - Na <sup>+</sup> D. None of these
10	Select the major product obtained from the addition HBr to 1-methyl cyclohexene.	A. 1- bromo -2- methyl cyclohexane B. 6- bromo-1- methyl cyclohex - 1- ene C. 3- bromo-1- methyl cyclohex -1- ene D. 1- bromo-1- methyl cyclohexane
11	Which of the following reagent cannot be used to detect the phenolic group.	A. Neutral FeCl <sub>3</sub> B. I <sub>2</sub> /NaOH C. NaOH solution D. Br <sub>2</sub> /H <sub>2</sub> O
12	Sugar and common salt in a mixture can be separated through then process of.	A. Sublimation B. Distillation C. Ion exchange D. Crystallization from solution in ethanol
13	Monomer of neoprene rubber to	A. 1-chloro 1,3- butadiene B. 2- chlro, 1,3-butadiene C. 2-Bromo -1,3- butadiene D. 2-Methyl 1,3-butadiene
14	Given A + 3B _____ 2C + D This reaction is first under with respect to reactant A and second order with respect to reactant B . If the concentration of A is doubled and the concentration of B is	A. Increase ,2 B. Decrease ,2 C. Increase ,4

halved, the rate of the reaction would \_\_\_\_\_ by factor of \_\_\_\_\_

- C. increase ,4
- D. Decrease ,4

- |    |  |  |
|----|--|--|
| 15 | Aromatic amine (X) was treated with alcoholic potash and another compound (Y) when foul smelling gas was formed with formula $C_2H_3N$ (Y) was formed by reacting a compound (Z) with $Cl_2$ in the presence of slaked lime . The compound (Z) is  | A. $C_6H_5NC$<br>B. $CHCl_3$<br>C. $CH_3CH_2OH$<br>D. $C_6H_5NH_2$   |
| 16 | An organic liquid (X) containing C, H and H has a pleasant odour with a boiling point of 78 oC. On boiling X with conc. $H_2SO_4$ a colourless gas is produced which decolourless bromine water and alkaline $KMnO_4$ One mole of this gas also takes one mole of $H_2$ . The organic liquid (X) is. | A. n- $C_3H_7OH$<br>B. iso- $C_3H_7OH$<br>C. $C_2H_5CHO$<br>D. $CH_3CH_2OH$  |
| 17 | Process of separating the racemic mixture into optically active isomers is known as.   | A. Resolution<br>B. Racemisation<br>C. Walden inversion<br>D. Epimerization  |
| 18 | Trimethylamine is a weaker base than dimethylamine is explained by   | A. Steric effect<br>B. Resonance effect<br>C. Inductive effect<br>D. All above   |
| 19 | Each of the following compound react with Grignard's reagent to form alkane except.  | A. Ethanal<br>B. Ethanoic acid<br>C. Ethanol<br>D. Ethyne  |
| 20 | Which of the following statement is not correct with respect to electrometric effect.  | A. It is permanent effect<br>B. It is brought into play instantaneously at the demand of attacking reagent<br>C. It proceeds a polar addition reaction<br>D. The original electronic condition is restored after the removal of tacking reagent. |