

PPSC Chemistry Part II Organic Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following reaction cannot be used for the synthesis of a amino acids.	A. Gabriel phthalimide B. Streckers synthesis C. Sorensen synthesis D. Schmidt synthesis
2	In the presence of dilute alkali monosaccharides undergo reversible isomerisation . The reaction known as.	A. Kiliani reaction B. Weermann rearrangement C. Lobry de Bruyn Van Ekenstein rearrangement D. Mutarotation
3	Glucose and fructose react with which of the following reagent to give same product.	A. Tollen's reagent B. Phenyl hydrazine C. Hydroxyl amine D. All of these
4	Which of the following is not a polysaccharide	A. Cellobiose B. Cellulose C. Insulin D. Amylase
5	Carbohydrates are characterized by the presence of.	A. Hydroxyl group B. Carbonyl group C. Asymmetric carbon D. All of these
6	Homolytic fission of covalent bond results in the formation of.	A. Free radicals B. Carbocations C. Carbanions D. Both B and C
7	Which of the following equations represent linear free energy relationship.	A. Hammett equation B. Taft equation C. Helmholtz equation D. Differential equation
8	Which of the following statement is not correct with respect to limitations of Hammett equation.	A. It is only applicable to aromatic systems B. Only applicable to aliphatic systems C. It is not valid for m-substituent
9	Which of the following statements is not correct with respect to applications of Hammett equations.	A. It develops a quantitative relationship between structure and reactivity B. This equation can be used to calculate the value of $pK_{sub>a</sub>}$ C. This equation does not help to calculate the rate of some reactions D. This equation has mechanistic implications
10	The equation which relates the reaction rates and equilibrium constants of many reactions is known as.	A. Taft equation B. Hammett equation C. Differential equation D. Linear equation
11	Which of the following haloacids is stronger acids.	A. FCH_2COOH B. $ClCH_2COOH$ C. $BrCH_2COOH$ D. ICH_2COOH
12	Which of the following groups exert -I effect.	A. $-NO_2$ B. $-CN$ C. $-COOH$ D. $\&C = O$
13	Which of the following factors effect the strengths of acids and bases.	A. Inductive effect B. Resonance effect C. Hydrogen effect D. All above
14	Which of the following statements not correct with the concept of Bronsted concept of acids	A. An acid can donate a proton B. A base can accept a proton

14	and bases.	<p>C. This concept has many bases that have OH⁻ ions</p> <p>D. This concept is more general</p>
15	Which of the following statement is not correct in respect of Arrhenius concept.	<p>A. This concept is applicable only for aqueous systems.</p> <p>B. Neutralization takes place in aqueous medium only</p> <p>C. H⁺ ion concept remain as such in water</p> <p>D. This concept is applicable for non aqueous system only.</p>
16	Inorganic acids (HCl, HBr, HNO ₃ etc) have K value.	<p>A. < 1</p> <p>B. > 1</p> <p>C. > 10</p> <p>D. < 10</p>
17	Which of the following statements do not represent Lewis idea of acids and base?	<p>A. Compounds which have completely filled orbitals</p> <p>B. Compounds which have incompletely filled orbitals</p> <p>C. Compounds in which the central atom can expand its octet</p> <p>D. All simple metal ions like Ag⁺, Al³⁺ etc.</p>
18	Which of the following does not represent Lewis base.	<p>A. Pyridine</p> <p>B. NaNH₂</p> <p>C. PCl₃</p> <p>D. NaOH</p>
19	Which of the following does not represent Lewis acid.	<p>A. ZnCl₂</p> <p>B. FeCl₂</p> <p>C. BF₃</p> <p>D. BuLi</p>
20	Which of the following pairs does not represent Lowery acid base pair.	<p>A. H₂O + NH₃</p> <p>B. H₂O + H₂O</p> <p>C. HCl + H₂O</p> <p>D. CH₃NH₂ + BF₃</p>