PPSC Chemistry Part I Physical Chemistry Online Test

Sr	Questions	Answers Choice
1	The depolarizer used in dry cell batteries in.	A. NH4Cl B. MnO2 C. KOH D. Na2PO4
2	The electrode Pt/Fe2+ (C1) Fe 3+ (C2) belong to the type.	A. Gas electrodes B. Inert metal electrodes C. Magam electrodes D. Metal metal insoluble salt electrode
3	Which of the following statement is not correct regarding galvanic cells.	A. Oxidation occurs at the anode B. Ions carry current inside the cell C. Electrons flow around the external circuit. form cathode to anode D. When the e.m.f. of the cell is positive cell reactionis spontaneous
4	In which of the following compound valency of carbon in 4 but its oxidation number is zero	A. Methane B. Carbon dioxide C. Carbon monoxide D. Formaldeyde
5	The oxidation Number of I in HIO4 is.	A. +6 B. + 7 C. + 3 D. + 14
6	During reaction of copper with aqueous solution of silver nitrate	A. Silver atoms are reduced B. Cu2+ ions are reduced C. Silver ions are reduced D. No³ ions are reduced
7	Which of the following process always involve the decrease in oxidation number.	A. Hydrolysis B. Elecomposition C. Oxidation D. Reduction
8	The equivalent conductance (^) and molar conductance (^ m) of BaSO4 are related as.	A. ^ = ^m/2 B. ^/2 = ^m C. ^ = ^m D. ^ = ^m/4
9	The equivalent conductance of a 1 N solution of an electrolyte is nearly	A. The same as its specific conductance B. 10³ times more than its specific conductance C. 10-3 times its specific conductance D. 100 times its specific conductance.
10	According to the Debye-Huckel theory of strong electrolytes, and ion moving in an atmosphere of oppositely charged ions experience a drag This effect is known as	A. Aaymmetric effect B. Electrophoretic effect C. Inter ionic effect D. Concentration effect
11	During the titration of weak acid against NaOH the conductance of the solution after the neutralization point.	A. Is constant B. Decreases C. Varies irregularly D. Increase
12	When a concentrated solute of an electrolyte is diluted.	A. Its specific conductance increases B. Its equivalent conductance decreases C. The specific conductance decreases and equivalent conductance increases D. Both specific and equivalent conductance increase
13	If for a solution of an electrolyte. It is the transport number of the cation, then the transport number of the anion I, is equal to	A. t/2 B. 1 - t+ C. 1 + t+ D. (i -t) /2

14	The fraction of the total current carried to an ion is called itss.	A. Ionic mobility B. Transport number C. Limiting ionic conductance D. None of these
15	Equivalent conductance is expressed in the units.	A. S cm-1 eq -1 B. S cm eq-1 C. S cm2 eq-1 D. S cm2 eq
16	Which of the following statements is not correct.	A. The conductance of one cm3 of a material is called specific conductance B. Specific conductance increase while equivalent conductance decreases on progressive dilution C. The limiting equivalent conductance of weak electrolytes cannot be determined by extrapolation of the plot of A against concentration D. The conductivity of metals is due to the movement of elctrons.
17	Which of the following statement is not correct with reference to cell constant.	A. The dimensions of cell constant is cm-1 B. It is used to determine the specific conductance C. It is measured with KCl solution D. Specific conductance does not vary with concentration.
18	Which of the following expressions represent the equivalent conductance.	A. A = l s x 1000/V B. A = Ls x 1000/C C. A = Ls l/A D. A = Ls/V
19	The conductance of 1 cm3 of an electrolytes solution is called its.	A. Specific resistance B. Specific conductance C. Molar conductance D. Equivalent conductance
20	The unit of specific conductance will be	A. S cm-1 B. Ohm cm C. Ohm cm-1 D. Mho cm