PPSC Chemistry Part I Physical Chemistry Online Test

Sr	Questions	Answers Choice
1	The tyndall effect was used by Zsigmondy to device.	A. The ultramicroscope B. The ultracentrifuge C. The osmometer D. Electrodialysis
2	A colloidal system in which a liquid is dispersed in a solid is called a/an	A. Emulsion B. Sol C. Gel D. Precipitate
3	An emulsifier is an agents which	A. Stabilizes an emulsion B. Homgeneises and emulsion C. Causes coagulation of an emulsion D. Helps in the formation of an emulsion
4	A silver iodide and was prepared by mixing KI and AgNO2 solution with the AgNO2 in slight excess. Which of the following descriptions is correct regarding is not particles.	A. Negatively charged because of the excess of NO-3 ions B. Positively charged because of the excess of Ag+ ions in the AgI lattice C. Negatively charged because I ions are adsorbed from the KI solution D. Neutral
5	Which of the following will be most effective int he coagulntion of Fe (OH)3 sol.	A. NaCl B. MgSO4 C. AlCl3 D. Mg3 (PO4)2
6	The stabilization of the dispersed phase in a lyophobic sol is due to	A. Liking for the dispersion medium B. The surface tension of the medium C. The formation of an electrical layer between the two phases D. The viscosity of the medium
7	Which of the following electrolytes will be most effective in the coagulation of arsenic sulphide sol.	A. NaNO3 B. Al PO4 C. MgSO4 D. K4[Fe(CN)6]
8	When a strong beam of light is passed through a colloidal solution, the light will	A. Be reflected B. Be scattered C. Pass unchanged D. Be dispersed
9	In the process of electrosmosis	A. Colloidial particles move towards the electrodes B. Both colloidal particles and dis persons medium move C. Only dispersion medium moves to carry the current D. Positively charged colloidal particles move, but negatively charged particles remain stationary
10	Smoke is a dispersion of	A. Gas in gas B. Gas in solid C. Solid in gas D. Liquid in gas
11	The migration of positively charged colloidal particles, under an electrical field , towards the cathode is called.	A. Cataphoresis B. Electroamosis C. Sedimentation D. Electrodialysis
12	The process of removing disolved impurities from a colloidal system, by means of diffusion through a suitable membrane under the influence of an electric field , is called.	A. Electrosmosis B. Electrodialysis C. Electrophoresis D. Peptization
13	Which of the following colligative properties can be used to characterize colloidal particles.	A. Lowering in vapour pressure B. Elevation in boiling point C. Depression in freezing point D. Osmotic pressure

14	Colloids can be purified by	A. Peptization B. Coagulation C. The Breeding are method D. Dialysis
15	Which of the following is the cause of Brownian movement of colloidal particles.	A. Convection currents int he fluid B. Bombardment by the molecules of the dispersion medium C. Setting of dispersed phase under gravity. D. Thermal gradient in the medium
16	Which of the following statement is false regarding lyphilic sols.	A. The colloidal particles show a linking for the dispersion medium B. These are generally easy to prepare C. These are more stable than lyophobic sols D. The stability of the sols is mainly due to the electrical double layer
17	A system is said to be in the colloidal state if the particle size of the dispersed phase ranges from	A. _o l to 10 A B. 10 to 10000 A C. 10 to 100 A D. 1000 to 10000 A
18	In a standard Weaton cell the cathode is	A. Cadmium amalgam B. Mercury C. Platinum D. Carbon
19	Concentration polarization arises because of the	A. Different concentrations of solutions in the two half cells B. Changes in the concentration of electrolyte around the electrode from bulk concentration C. Reversible nature of the cell D. Variation in temperature during measurements
20	A half cell reaction is one that	A. Occurs at one electrode B. Goes only half way to completion C. Involves a half mole of the concentration of the solution D. Always oxidizes