PPSC Chemistry Part I Physical Chemistry Online Test

Sr	Questions	Answers Choice
1	30 mL of an acid solution is neutralized by 15 mL of 0.2 N base. The strength of acid solution is.	A. 0.1 N B. 0.15 N C. 0.3 N D. 0.4 N
2	The normality of 2.3 M H2SO4 solution is.	A. 0.46 N B. 0.23 N C. 2.3 N D. 4.6 N
3	Which of the following solution has highest normality.	A. 1 N H2PO4 B. 0.5 N H2SO4 C. 6 g NaOH per 100 cm3 D. 4 g NaOH PER 1000 cm3
4	Which of the following expression is correct.	A. C = n/RT B. C = RT/n C. RT = Cn D. Cn = 1/RT
5	At the some temperature 0.1 M solution of urea is isotonic with.	A. 0.1 M glucose solution B. 0.1 M NaCl solution C. 0.1 M urea solution D. 0.1 M BaCl2 solution
6	Sea water is converted into fresh water bases upon the phenomenon of.	A. Plasmolysis B. Sedimentation C. Diffusion D. Osmosis E. Reverse osmosis
7	The flow of solvent into a solution when two are separated by a semi -permeable membrane is called.	A. Mixing B. Effusion C. Diffusion D. Osmosis
8	Which of the following solution would have the largest depression in freezing point.	A. 1% glucose B. 1 % KCl C. 1 % AlCl3 D. 1 % BaCl2
9	Which of the following solution would exhibit abnormal colligative proportions.	A. 0.1 M NaCl B. 0.1 M urea C. 0.1 M sucrose D. 0.1 M glucose
10	The osmotic pressure of a solution with definite composition.	A. Varies directly as the volume and temperature. B. Various inversely as the temperature. C. Varies inversely as the volume and directly as the temperature. D. None of the above
11	The relative lowering of vapour pressure of a solution on the addition of non -volatile solute.	A. Is equal to the mole fraction of solute B. Is equal to the sum of the mole fraction of the solute and solvent C. Depends upon the nature of the solute D. Depends upon the mole fraction of the solvent
12	Iso-osmotic solutions are those which have the same.	A. Vapour pressure lowering B. Osmotic pressure C. Molality D. Boiling point elevation
13	The freezing point of a solvent	A. Will increase on adding a solute B. Will decrease on adding a solute C. Will note change on adding solute D. None of the above
14	The temperature at which the vapour pressure becomes equal to external pressure is called.	A. Saturation point B. Critical temperature

		C. Consolute temperature D. Boiling point
15	Which of the following properties does not depend upon the numeb rof solute particles.	A. Elevation in B.P. B. Osmotic pressure C. Depression in F.P. D. Boiling point of the solvent
16	Which of the following is not a colligative property.	A. Elevation of B.P. B. Depresaion in F.P C. Viscosity D. Osmotic pressure
17	For dilute solutions colligative properties depend on.	A. The number of the particles of the solute and nature of solvent. B. The number of the solute particles and on their nature C. The number of the solute particles and nature of solute and solvent D. The number of the solute particles and irrespective of the nature of the solute and solvent.
18	According to Henry's Law. the mole fraction of a gas (x) dissolved in a solvent is related to the pressure of the gas.	A. x = k/p B. x = p/k C. x = k D. p = k/x
19	The law which relates the solubility of a gas to its pressure is called.	A. Raoult's law B. Nernst law C. Ostwald's law D. Hennery's law
20	The temperature at which two conjugate solutions change into one homogeneous solution is called.	A. Azeotrope B. Conjugate temperature C. Consolute temperature D. Transition temperature