PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	The hybridization of S in SO2 is.	A. sp B. sp2 C. sp3 D. dsp2
2	B.P of heavy water is	A. equal to that of ordinary water B. greater than that of ordinary water C. Less than that of ordinary water D. equal to that of distilled water
3	When a lead a storage battery is discharged .	A. SO2 is evolved B. PbS is consumed C. Pb is formed D. H2SO4 is consumed
4	Pick out the incorrect statement for SO2	A. It turns filter paper moistened with acidified K2Cr2O7 B. It turns starch iodate paper blue C. It does not react with chlorine in presence of charcoal D. It decolourises acidified KMnO4 solution.
5	Oxalic acid when heated withe conc. H2SO4 it gives out.	A. H2O and CO2 B. CO and CO2 C. CO2 and H2S D. Oxalic sulphate
6	Which among the following is a Talse statement.	A. SO3 is obtained by the catalytic oxidation of SO2 B. SO3 has trigonal planar geometry in gaseous state C. SO3 in nauseous state has all S-O bonds equivalent D. SO3 gas shows more solubility in water than in H2SO4
7	What is the following is incorrect.	A. Water is more polar than H2S B. H2O2 is a planar molecule C. Heavy water is produced by the exhauative electrolysis of water made acidic D. H2O2 act both as oxidising as well as reducing agent in acidic medium
8	Pick out the incorrect statement regarding ozone.	A. O3 is an unstbale dark blue diamagnetic gas B. The central oxygen in O3 is sp3 hybridized C. It cause the tailing of mercury D. It does not react with KOH
9	Which of the following is not true of ozone.	A. It is a strong electilizing agent B. It attacks organic compounds containing carbon carbon double bond C. Its molecular is linear and has two different O-O bond lengths D. It is more powerful oxidising agent at molecular oxygen
10	Which of the following reactions is employed to produce ozone in the laboratory.	A. Exposure of air to UV light B. Reaction of F2 with H2O at low temperature C. Reaction SO2 with H2O2 D. Passage of silent electric discharge through oxygen
11	O2 molecule is.	A. Fermagnetic B. Forromagnetic C. Paramagnetic D. Diamagnetic
12	Black and white photographic film contain small grains of.	A. Silver bromide B. Silver cholride

		C. Silver iodide D. Any of above
13	Perdisulphuric acid is.	A. Marshal acid B. Caro acid C. None of above D. Any of above
14	H2SO4 is used	A. In the preparation of aqua regia B. In the purification of gold and silver C. In the dental filling D. None of above
15	The specific gravity of H2SO4 is	A. 1.37 B. 1.84 C. 1.17 D. 1.57
16	The contact process is mainly used when acid is required for the manufactures of.	A. Explosives B. Fine chemicals C. Lead accumulators D. All above
17	The H2SO4 obtained by the contact process having purity	A. 70% B. 74% C. 78% D. 82%
18	Gases and dust particles are removed from H2SO4 by	A. Tydal effect B. Drying tower C. Absorption tower D. Contact converter
19	The commonly used catalyst in the manufacture of H2SO4	A. Fe2O3 with a little CuO B. v2O5 C. Platinized asbestos and MgSO4 D. All above
20	H2SO4 is manufactured by	A. The lead chamber process B. The contact process C. Both A and B D. The Ostwald's process
21	H2SO4 acts as gent	A. Reducing B. Oxidizing C. Both A and B D. None of above
22	A thionic acid	A. H2S2O3 B. H2S2O6 C. H2S2O8 D. H2S2O7
23	Is a peroxy acid	A. H2SO5 B. H2S2O6 C. H2SO4 D. H2S2O7
24	SO3 exists in form	A. a -so3 B. b-SO3 C. gama SO3 D. All above
25	The structure of SO2 is	A. Linear B. Angular C. V-shaped D. Planner
26	SO2 acts as	A. Lewis base B. Lewis acid C. Botha A and B D. None of above
27	The formula of sulphur sesquioxide	A. SO4 B. S2O7 C. S2O3 D. SO3
28	Molecule of oxygen is	A. Diamagnetic B. Paramagnetic C. Both A and B D. None of above
29	Oxygen and sulphur exist in state	A. Free B. Combined C. _{Both free & combined} D. None of above

		D. None of above
30	Electronegativity of oxygen is.	A. 2,5 B. 3,5 C. 2,4 D. 2.1