

PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	The hybridization of S in SO2 is.	A. sp B. sp2 C. sp3 D. dsp2
2	B.P of heavy water is	A. equal to that of ordinary water B. greater than that of ordinary water C. Less than that of ordinary water D. equal to that of distilled water
3	When a lead a storage battery is discharged .	A. SO2 is evolved B. PbS is consumed C. Pb is formed D. H2SO4 is consumed
4	Pick out the incorrect statement for SO2	A. It turns filter paper moistened with acidified K2Cr2O7 B. It turns starch iodate paper blue C. It does not react with chlorine in presence of charcoal D. It decolourises acidified KMnO4 solution.
5	Oxalic acid when heated withe conc. H2SO4 it gives out.	A. H2O and CO2 B. CO and CO2 C. CO2 and H2S D. Oxalic sulphate
6	Which among the following is a Talse statement.	A. SO3 is obtained by the catalytic oxidation of SO2 B. SO3 has trigonal planar geometry in gaseous state C. SO3 in nauseous state has all S-O bonds equivalent D. SO3 gas shows more solubility in water than in H2SO4
7	What is the following is incorrect.	A. Water is more polar than H2S B. H2O2 is a planar molecule C. Heavy water is produced by the exhauative electrolysis of water made acidic D. H2O2 act both as oxidising as well as reducing agent in acidic medium
8	Pick out the incorrect statement regarding ozone.	A. O3 is an unstbale dark blue diamagnetic gas B. The central oxygen in O3 is sp3 hybridized C. It cause the tailing of mercury D. It does not react with KOH
9	Which of the following is not true of ozone.	A. It is a strong electilizing agent B. It attacks organic compounds containing carbon carbon double bond C. Its molecular is linear and has two different O-O bond lengths D. It is more powerful oxidising agent at molecular oxygen
10	Which of the following reactions is employed to produce ozone in the laboratory.	A. Exposure of air to UV light B. Reaction of F2 with H2O at low temperature C. Reaction SO2 with H2O2 D. Passage of silent electric discharge through oxygen
11	O2 molecule is.	A. Fermagnetic B. Forromagnetic C. Paramagnetic D. Diamagnetic
12	Black and white photographic film contain small grains of.	A. Silver bromide B. Silver cholride

B. In the purification of gold and silver C. In the dental filling D. None of above A. 1.37 B. 1.84 C. 1.17 D. 1.57 The contact process is mainly used when acid is required for the manufactures of. The H2SO4 obtained by the contact process having purity A. 70% B. 74% C. 78% D. 82% A. Tydal effect B. Drying tower C. Absorption tower D. Contact converter A. Fe2O3 with a little CuO B. V2O5			C. Silver logide D. Any of above
B. In the purification of gold and solver C. In the dental filling C. In the specific gravity of H2SC4 is C. In the contact process is mainly used when acid is required for the manufactures of. The contact process is mainly used when acid is required for the manufactures of. A Explosives E. Fine chemicals C. Lead accountablers C. Lead accountablers C. Lead accountablers C. Lead A. Tystal effect E. C. In the L2SC4 obtained by the contact process having purity C. In the L2SC4 C. I	13	Perdisulphuric acid is.	B. Caro acid C. None of above
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The contact process is mainly used when acid is required for the manufactures of. Contact process is mainly used when acid is required for the manufactures of.	15	The specific gravity of H2SO4 is	B. 1.84 C. 1.17
17 The H2SO4 obtained by the contact process having purity 18 C 78% D. 82% 18 Gases and dust particles are removed from H2SO4 by 19 The commonly used catalyst in the manufacture of H2SO4 20 H2SO4 is manufactured by 21 H2SO4 acts as gent 22 A thionic acid 23 Is a peroxy acid 24 SO3 exists in form 25 The structure of SO2 is 26 SO2 acts as 27 The formula of sulphur sesquioxide 28 Molecule of oxygen is 28 Molecule of oxygen is 29 Molecule of oxygen is 20 Massocia A Demangeric D. Son	16	The contact process is mainly used when acid is required for the manufactures of.	B. Fine chemicals C. Lead accumulators
B. Drying tower C. Absorption tower C. Absorption tower D. Contact converter 19 The commonly used catalyst in the manufacture of H2SO4 20 H2SO4 is manufactured by 20 H2SO4 acts as gent 21 H2SO4 acts as gent 22 A thionic acid 23 Is a peroxy acid 24 SO3 exists in form 25 The structure of SO2 is 26 SO2 acts as 27 The formula of sulphur sesquioxide 28 Molecule of oxygen is 29 Molecule of oxygen is 20 Molecule of oxygen is 20 Regards 21 H2SO4 acts as gent 21 H2SO4 acts as gent 22 A thionic acid 23 Is a peroxy acid 34 Reducing 35 Regards 36 Regards 37 Reducing 38 Regards	17	The H2SO4 obtained by the contact process having purity	B. 74% C. 78%
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23 Is a peroxy acid B H2S2O6 C. H2SO4 D. H2S2O7 A a -so3 B. b-SO3 C. gama SO3 D. All above 25 The structure of SO2 is 26 SO2 acts as 27 The formula of sulphur sesquioxide 28 Molecule of oxygen is B H2S2O6 C. H2SO4 D. H2SO7 C. spana SO3 D. All above A Linear B. Angular C. V-shaped D. Planner A Lewis base B. Lewis acid C. Botha A and B D. None of above A SO4 B. S2O7 C. S2O3 D. SO3 A Diamagnetic B. Paramagnetic C. Both A and B D. None of above A Free	22	A thionic acid	B. H2S2O6 C. H2S2O8
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28 Molecule of oxygen is B. Paramagnetic C. Both A and B D. None of above A. Free	27	The formula of sulphur sesquioxide	A. SO4 B. S2O7 C. S2O3
	28	Molecule of oxygen is	B. Paramagnetic C. Both A and B
29 Oxygen and sulphur exist in state C. _{Both free & amp; combined} D. None of above	29	Oxygen and sulphur exist in state	B. Combined C. _{Both free & amp; combined}

C. Silver logide

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30 Electronegativity of oxygen is. A. 2,5 B. 3,5 C. 2,4 D. 2.1