PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Un-like s -block elements d-block elements form which compounds as well	A. ionic compounds B. Co valent compound C. Co ordinate compounds D. None of above
2	Ten elements Sc (Z = 21) to Zn (Z = 30) fill their 4s orbitals first and then 3d orbitals are called elements. of.	A. 3 d series. B. 4d Series C. 5d Series D. None of above
3	Elements in which differentiating electron enters the (n-1) the d-orbitals of the (n-1) the main shell are called elements.	A. s- block B. p-block C. d-block D. f-block
4	Carbon in wrought iron is present as	A. Silicon carbide B. Iron carbide or comentite C. Graphite D. Partly as iron carbide and partly as graphite
5	Stainless steel contains.	A. Fe + Cr+ Ni B. Fe + Ni + Cu C. Fe + Cr+ Cu D. Cu + C + Ni
6	If Steel is heated to a temperature well below red heat and to then cooled slowly the process is called.	A. Annealing B. Quenching C. Tempering D. Nitriding
7	In extraction of iron, the furnace charge consists of iron ore, coke and limestone. The function of limestone is to act as.	A. An oxidation agent B. A reducing agent C. Flux D. Slag
8	Molten iron withdrawn from the blast furnace is called.	A. Wrought iron B. Pig iron C. Bessemer iron D. Stainless steel
9	In the forth floatation process for the purification of ores, the ore particles float because.	A. They are light B. Their surface is not easily vetted by water C. They bear electrostatic charge D. They are insoluble
10	The number used in cancer therapy is.	A. Fe B. Co C. Ni D. Rn
11	Transition elements, in general exhibit the following properties, except one Name that property.	A. Variable oxidation state B. Natural radioactivity C. Tendency to form complexes D. Formation of alloys
12	Which of the following statements false about transition metals.	A. They form complexes B. They show variable valency C. All transiting metal compounds are paramagnetic D. They form coloured ions
13	Which one of the following ions is colourless.	A. Cu+ B. Co2+ C. Ni2+ D. Fe3+
14	Gold dissolves in aqua regia forming	A. AuCl B. Au(NO3)3 C. AuCl3 D. HAuCl4
		A. AgF

15	Which of the following is soluble in water.	B. AgCl C. AgBr D. Ag I
16	Which of the following alloys contains Cu and Zn	A. Bronze B. Brass C. Gun metal D. Type metal
17	The rusting of iron is catalyzed by which of the following.	A. Fe B. H+ C. O2 D. Zn
18	Which of the following is not correct.	A. Rusting of iron can be stopped by increasing the concentration of CO2 in water B. Rusting of iron is electrochemical in nature. C. Rusting of iron takes place in moist air D. Rusting of iron produces hydrated iron (III) oxide
19	Finely divided iorn combines with CO to give	A. Fe(CO)3 B. Fe2 (CO)9 C. F33(CO)12 D. Fe(CO)6
20	In the metallurgy of iron, when limestone is added to the blast furnaced, the calcium ion ends up in	A. Slag B. Gangue C. Metallic calcium D. Calcium carbonate
21	Pick out incorrect statement about K2Cr2O7	A. It oxidizes acidified solution H2SO4 to S B. It oxidizes KI to I2 C. It oxidizes HCl to Cl2 D. It gives oxygen, when treated with cold conc. H2SO4
22	Pick out the incorrect statement about K2Cr2O7	A. It is thermally stable B. It dissolves in alkali to form chromate C. It oxidizes acidified FeSO4 solution to Fe2(SO4)3 D. It is used as cleansing agent for glassware. etc. When mixed with cold con. H2SO4
23	The yellow colour of chromates changes to orange red on acidification, due to the formation of.	A. Cr3+ B. Cr2O3 C. Cr2O7²⁻ D. Cro3
24	Which one of the following oxides is basic.	A. MnO B. Mn2O3 C. MnO2 D. Mn2O7
25	The first ionization energies of the elements of the first transition series. (Ti _____ Cu)	A. Increases as the atomic number increases B. decreases as the atomic number increases C. Do not show any change as the addition of electrons takes place in the inner (n-1) d -orbitals. D. Increases from Ti to Mn and then decreases from Mn to Cu
26	Which one of the following statements is not true.	A. Transition metals form alloys B. Transition metals form complexes C. Zn, Cd and Hg are transition metals D. K2[PtCl6] is a well known compound but corresponding nikel compound is not knonwn
27	Pick out the incorrect statements for transition metals.	A. They have low melting and boiling points B. 5d-element have higher energies than 3d or 4 d elements C. Zr and Hf have almost identical atomic and ionic radii D. They form interstitial compounds.
		A. Mn is a transition element while K+ is not B. [MnO4]is negatively charged while

28	The atomic number of Potassium is 19 and that of manganese is 25. Although the coloured of MnO4 is dark violet yet the K+ is colourles.. This is due to the fact that	B. [MnO4] is negatively charged while K+ has a positive charge C. The effective atomicnumebr of Mn is [MnO4] is 26 while for K+ the atomic number is 18 D. The Mn in a high positive oxidation state allows charge transfer transitions
29	Pick out the incorrect statement for transition metals.	A. Cu+ is not a transition metal ion B. Transition metals do not exhibit variable oxidation states C. Transition metal ions are coloured D. Transition metals and majority of their compounds are paramagnetic
30	The electronic configuration of chromium is 4s1, 3d5, The elements tungsten (W) belongs to the same group and has atomic number 74.The configuration of its valence shell is.	A. 5s1, 4d5 B. 6s1, 5d5 C. 6s1, 5d6 D. 6s1, 5d4