PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	The principal ores of copper are	A. Copper sulphides B. Copper oxides C. Both sulphides and oxides D. Copper carbonate
2	Copper occurs in nature as.	A. Native B. Combined C. Both native and combined D. None of the above
3	Which of the following is the second anciently known metal.	A. Nickel B. Copper C. Gold D. Silver
4	Carbon in wrought iron is present as	A. Silicon carbide B. Iron carbide cementite C. Graphite D. Partly as iron carbide and partly as graphite
5	Stainless steel contains	A. Fe+Cr+Ni B. Fe+Ni+Cu C. Fe + Cr+ Cu D. Cu + C + Ni
6	If steel is heated to a temperature well below red heated and is then cooled slowly the process is called.	A. Annealing B. Quenching C. Tempering D. Nitriding
7	In the extraction of iron, the furnace change consists of iron ore, coke and limestone. The function of limestone is to act as.	A. An oxidizing agent B. A reducing agent C. Flux D. Slag
8	Pig iron is also called.	A. Cast iron B. Steel C. Wrought iron D. Stainless steel
9	In the fourth flotation process for the purification of ores, the ore particles floats because.	A. They are light B. Their surface is not easily wetted by water C. They bear electrostatic charge D. They are insoluble
10	The most stable oxidation state of chromium is.	A. +6 B. +3 C. +4 D. +2
11	Zeigler Natta catalyst is.	A. Pt/PtO B. TiCl4/Al(C2H5)3 C. Pt/Rh D. Pt
12	Transition elements, in general, exhibit the following properties, except one, Name that property.	A. Variable oxidation state B. Natural radioactivity C. Tendency to form complexes D. Formation of alloys
13	Which of the following statements is false about transition metals.	A. They form complexes B. They show variable valency C. All transition metal compounds are paramagnetic D. They form coloured ions
14	Which one of the following ionsis colourless.	A. Cu+ B. Ni2+ C. Co2+ D. Fe3+
		A. Small sized metal ions

15	Colour in transition metal compounds in attributed to	B. Absorption of light in UV region C. Complete ns sub shell D. incomplete (n-1) sub shell
16	The maximum oxidation shown by managanese is.	A. +2 B. +7 C. +4 D. +5
17	Variable oxidation states is shown by	A. Normal eleemnts B. Metallic elements C. Non metallic elements D. Transition elements
18	Zinc oxide is.	A. A basic oxide B. An amphoteric oxide C. An acidic oxide D. A neutral oxide
19	The rusting of iron is catalysed by which of the following.	A. Fe B. O2 C. Zn D. H+
20	Finely divided iron combines with CO to give.	A. Fe(CO)5 B. Fe2(CO)9 C. Fe(CO)12 D. Fe(CO)6
21	In the metallurgy of iron, when limestone is added to the blast furnace, the calcium ion ends up in.	A. Slag B. Gangue C. Metallic calcium D. Calcium carbonate
22	Pick out incorrect statemtn about K2r2O7	A. It oxidizes acidified solution of H2S to S B. It oxidizes Ki TO i2 C. It oxidizes HCl to Cl2 D. It gives oxygen, when treated with cold conc. H2SO4
23	Pick out the incorrect statement about K2Cr2O7	A. It is thermally stable B. It dissolves in alkali to form chromate C. It oxidizes acidified FeSO4 solution to Fe2(SO4)3 D. It is used as cleansing agent for glassware, etc. when mixed wiht cold conc. H2SO4
24	The atomic number of potassium is 19 and that of mangness is 25 Although the colour of MnO4 is dark violet yet the K+ is colourless this is due to the fact that.	A. Mn is a transition element while K+ is not B. [MnO4]- is negatively charged while K+ has positive charge C. The effective atomic number of Mn is [MnO4]- is 26: while for K+ the atomic number is 18 D. The Mn is a high positive oxidation state allows charge transfer transitions.
25	The secondary valency of Conc. CoCl3. 6NH3.	A. 2 B. 4 C. 6 D. 8
26	If the absorbed light is green the transmitted light will be	A. Purple B. Orange C. Violet D. Black
27	[Ti(OH2)6]3+ gives colour	A. Green B. Red C. Purple D. Blue
28	The maximum absorption in [Ti(OH)2)6 3+ take place at wavelength of.	A. 4000 A^o B. 5000 A^o C. 6000 A^o D. 10000 A^o
29	The solution of the transition metal complexes having one or more unpaired electrons in the d-orbital are.	A. Coloured B. Colourless C. White D. None of above
30	When metal orbital are rotated in octahedral field the following representation obtained.	A. t2 g + lg B. a1 g C. t 1 u

D. All above