PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Which of the following salt is colourless.	A. Zn salt B. Co salt C. Ni salt D. Mn salt
2	Which of the following sulphide is yellow in colour.	A. HgS B. PbS C. CdS D. SnS
3	Which of the following chloride is soluble in hot water.	A. Hg2Cl2 B. AgCl C. PbCl2 D. All above
4	NH4OH in the presence of H2S is used as a group reagent for which of the following group.	A. Group I B. Group II C. Group III D. Group IV
5	Which of the following group reagent is used for III group of basic radical.	A. Dilute HCl B. H2S + HCl C. NH4OH + NH4Cl D. NH4OH + H2S
6	Which of the following radical is a member of VI group.	A. Mg²⁺ B. Na+ C. K+ D. NH₄⁺ E. All above
7	Which of the following radical is not a member of IV group.	A. Mg²⁺ B. Co²⁺ C. Ni²⁺ D. Mn²⁺
8	Which of the following radical is not a member of III group	A. Al³⁺ B. Fe²⁺ C. Ca²⁺ D. Fe³⁺
9	Which of the following radical is not a member of II group.	A. Cu²⁺ B. Cd²⁺ C. Ba³⁺ D. K+
10	Which of the following is not an acid radical	A. Cl- B. Br- C. K+ D. I-
11	Which of the following species is not a basic radical.	A. Ag+ B. Cl- C. Ba2+ D. K+
12	Which of the following is not a physical test.	A. Colour test B. Flame test C. Beed test D. Wet test
13	Yellow green flame is observed with	A. Calcium salt B. Barium salt C. Strontium salt D. Sodium salt
14	Dull red flame is observed with	A. Calcium salt B. Barium salt C. Strontium salt D. Sodium salt
15	Yellow colour of the flame is observed with	A. Calcium salt B. Barium salt C. Sodium salt

		D. Potassium salt
16	Which of the following methods is used in qualitativ eanalysis.	A. Physical method B. Chemical method C. Instrumental method D. All above
17	Which of the following combination is used to make buffer.	A. NaOH and HCl B. KOH and H2SO4 C. CH3COOH and CH3COONa D. CH3COOH and NH4OH
18	The pH of 0.01 N NaOH is.	A. 12 B. 13 C. 14 D. 11
19	The pH of 0.001 N HCl is	A. 1 B. 2 C. 3 D. 4
20	It has been observed that if one goes on adding KNO3 solution to a precipitate of AgCl the solubility of these precipitates goes on increasing with increasing concentration of K+ and NO3= ions which are not common to AgCl This is due to which effect.	A. Divers ion effect B. Uncommon ion effect C. Activity effect D. All above
21	It is known that AgCl is insoluble in HNO3 but dissolves readily in NH4OH solution .Which of the following statement is not correct.	A. Ag ion reacts to form complex with NH4OH solution B. The concentration of Ag ion decreases C. Ionic product is less than the solubility product D. Ionic product is greater than solubility product
22	When to a solution of weak electrolyte a strong electrolyte with a common ion is added, the dissociation of weak electrolytes is suppressed . This is known as.	A. Stark effect B. Salt effect C. Common ion effect D. Zeman effect
23	In second group of inorganic qualitative analysis, the S^{2-} ions does not form precipitate with which of the following ions.	A. ^{Hg2+} B. Cu²⁺ C. Ai³⁺ D. Cd³⁺
24	Which of the following statement is not correct regarding dissociation constant (Ka)?	A. It is a measure of the tendency of an acid to split up into ions B. The greater the value of Ka, more is the dissociation C. It is determined by conductimetric method D. It is not a proper parameter for weak acids
25	Which of the following is not strong electrolytes.	A. HCl B. H2SO4 C. HNO3 D. CH3COOH
26	Which of the following substance is not weak electrolyte.	A. CH3COOH B. NH4OH C. Oxalic Acid D. NaCl
27	In order to increase the rate of the reaction one should.	A. Increase the concentration of products B. Decrease the concentration of reactants C. Decreases the concentration of products D. Both C and D statement are correct
28	If a chemical reaction in equilibrium is subjected to a change the reaction tends to more in such a direction that the effect of the change would be neutralized This is a statement of.	A. Law of mass action B. Le Chatlier's principle C. Henery's law D. Correspondence principle
29	The equilibrium constant value for a chemical reaction is 5×10^{20} which of the following statement is true with respect to this value.	A. Reaction will be reversible B. Reaction will proceed in backward direction C. Reaction is at equilibrium D. Reaction will proceed in the forward direction
30	The rate of a chemical reaction is proportional to the product of the active mean of the reactants, This is a statement of.	A. Law of dynamic equilibrium B. Le Chatlier's principle C. Law of mass action D. Solubility product principle

D. Solubility product principle