PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Which of the following is not an adsorption indicator.	A. Eosin B. Bromocrsol green C. Fluorescein D. Phenolphthalein
2	Which of the following is not an organic precipitating agent.	A. Diemethglyoxime B. Cuperon C. Oxime D. Acetate
3	Which of the following anionic species is not separated by gravimetric analysis.	A. Cl- B. SO4 C. CH3COO- D. PO4
4	Which of the following metal ion cannot be catimated by gravimetric analysis.	A. K+ B. Ca²⁺ C. Al3+ D. Zn2+
5	Which of the following analytical techniques can be used to extract metal ion chelates.	A. Solvent extractions B. Evaporation C. GC D. Distillation
6	Which of the following is not a ligand or complexing agent.	A. NH3 B. CH3COOH C. EDTA D. CN-
7	The number of bonds formed by the central atom is called its.	A. Valence number B. Complex number C. Coordination number D. Avogadro's number
8	Which of the following species is determined by complex metric titrations.	A. K+ B. Na+ C. Ca+ D. Cl-
9	Complexing reactions are useful for which of the following method of analysis	A. Gravimetry B. Spectrophotometry C. Interfering ions masking D. All of the above
10	Amino acids are are important in biochemistry which of the following statements is not correct regarding amino acids.	A. These are amphoteric substances tend to undergo internal protein transfer B. In aqueous solutions these substances tend to undergo internal proton transfer C. These for zwitter ion in aqueous medium D. These always contain two amino groups.
11	Which of the following is the best indicator for titration of NH4OH with HCl.	A. Methyl red B. Methyl orange C. Eosin D. Phenolphthalein
12	Which of the following is the best indicator for titration of CH3COH with NaOH	A. Methyl orange B. Methyl red C. Phenophthalein D. Eosin
13	When CH3COOH is titrated against NaOH the pH as the equivalence point is.	A. 7 B. > 7 C. o >7 D. 6.8
14	When HCl is titrated against NaOH , the pH at the equivalence point is.	A. zero B. > 7 C. < 7

		D. 14
15	An indicator for an acid base titration is a	A. Weak acid B. Weak base C. Strong acid D. Strong base E. Both A and B
16	The point at which the reaction is observed to be complete is called.	A. The equivalence point B. The end point C. The triplet point D. The equilibrium point
17	An acid base titration involves a neutralization reaction in which an acid is reacted with an equivalent amount of base The titrant is always a strong acid or base The analyte may be	A. Strong acid B. Strong base C. Weak acid and Weak base D. All above
18	Which of the following is not a component of HPLC system.	A. Pumps B. Columns C. Particle collector D. Injection system.
19	Which of the following functional groups is not involved in ion exchange chromatography.	A. Weak acids B. Strong acids C. Strong bases D. Carbohydrates
20	Which of the following statement is not related with the advantages of TLC.	A. A variety of adsorbents can be used B. The thickness of adsorbent can be varied C. Fluorescence can be introduced D. Different detectors can be used
21	Which of the following materials is not suitable as adsorbent for chromatography.	A. Silica gel B. Activated charocal C. Alumina D. Calciu7m chloride
22	Which of the following techniques is capable of seperating minute quantities of the substances in a relatively short times with high resolutions.	A. Gel electrophoresis B. Capillary electrophoresis C. GC D. HPLC
23	Which of the following technique is used to separate substance based on their charge to mass ratio.	A. HPLC B. HPTLC C. GC D. Electrophoresis
24	Which of the following detector is used in HPLC system.	A. Differential refractometer detector B. UV detector C. Diode array detector D. All above
25	Which of the following detector is used for compounds containing electronegative atoms.	A. Mass specdtrometer B. ECD C. TCD D. UV-detector
26	Which of the following detector is used in GC analysis	A. Thermal conductivity detector B. Flame ionization detector C. Mass spectrometer D. All above
27	A well packed column may hve	A. 100 plates /m B. 1000 plates /m C. 10 plates /m D. 10,000 plates/m
28	Which of the following information is correct about a typical packed column in GC.	A. 10-100 m long and 2 to 6 cm to diameter B. 1-10 m long and 0.2 to 0.6 cm in diameter C. 0.1-1 m long and 0.02 to 0.00 cm in diameter D. None of the above
29	Which of the following gas is not used as carrier gas in GC.	A. Argon B. Nitrogen C. Helium D. CO2
30	Which of the following is not a component of a gas chromatography system.	A. Carrier gas B. Capillary column C. Packed column D. Cathode lamp