

PPSC Chemistry Full Book Test

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Sr	Questions	Answers Choice
1	In which of the following techniques the solvated molecules are separated according to their size by their ability to penetrate a sieve like structure.	A. Adsorption chromatography B. Partition chromatography C. Ion exchange chromatography D. Gel permeation chromatography
2	It is possible to distinguish between optical isomers.	A. Using chemical tests B. By mass spectrometry C. By IR spectroscopy D. By polarimetry
3	The criteria for aromatically is presence of	A. Uneaturations B. Cyclic structure C. Presence of 4nx electrons D. Presence of 4n + 2n electrons
4	Which number of halogen family does not show positive oxidation state.	A. Fluorine B. Chlorine C. Bromine D. lodine
5	Which of the following pollutants does not a leave a residue.	A. Air pollutant B. Chemical pollutant C. Soil pollutant D. Noise pollutant
6	Beryllium shows diagonal relationship with.	A. Mg B. Al C. Na D. B
7	The Lewis formula of SOCl2, the total number of bond pairs and lone pairs of electron around sulphur are.	A. 2 ,1 B. 2,2 C. 3,1 D. 3,0
8	An indicator for an acid base titration is a	A. Weak acid B. Weak base C. Strong acid D. Strong base E. Both A and B
9	The dye which is a constituent of Skiffs reagent used for detection formaldehyde group is.	A. Gentain violet B. Megneta C. Phenolphthalein D. Rosolic acid
10	Which of the following statements regarding phenols is not correct.	A. Phenol Is are stronger acids than water and alcohols. B. Phenol are weaker acids than carboxylic acids C. Phenol are solubel in both aqueous NaOH and aqueous sodium hydrogen carbonate D. Phenoxides ions are more stable than the corresponding phenol
11	The compound (CH3)3COH according to IUPAC is known as.	A. Tert Butanol B. 2,2 -Dumethyl -Propanol C. 2- Methyl -2-propanol D. Tert Alcohol
12	Which of the following statements is not a part of Bohr's theory of the hydrogen atom.	A. An electron in an atom revolves aroung the nucleus only in circular paths. B. An electron does not absorb energy in the stationary orbit C. An electron does not emit energy in the stationary orbit D. Energy is emitted or absorbed in a discrete amount from the stationary orbit
13	The geometry of Xe F2 is	A. Triangular planar B. Square planar C. Linear

		D. Trigonal bipyramidal
14	Enzymatic action is heat at a fixed	A. Temparature B. pH C. Both of these D. None of these
15	The number of moles of solute dissolved in 1000 gram of the solvent is called	A. Formality B. Molality C. Molarity D. Mole fraction
16	Which one of the following has the highest boiling point.	A. H2O B. H2S C. H2Se D. H2Te
17	According to Henry's Law. the mole fraction of a gas (x) dissolved in a solvent is related to the pressure of the gas.	A. x = k/p B. x = p/k C. x = k D. p = k/x
18	Thermogravimetic analysis has application is which of the following fields	A. Gravimetric analysis B. Discovery of new methods ofseparation C. Determination of purity and thermal stability D. All above
19	The strongest acid is.	A. HNO2 B. HNO3 C. H2N2O2 D. HNOS
20	The solution of the transition metal complexes having one or more unpaired electrons in the d-orbital are.	A. Coloured B. Colourless C. White D. None of above