# PPSC Chemistry Full Book Test

| Sr | Questions | Answers Choice |
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| 1 | Which of the following is the major process when neopentyl bromide is dehydrogenate with alcoholic potash. | A. 2- methyl -1- butene<br>B. 2- methyl- 1- butene<br>C. 2,3 -dimethyl butene<br>D. 2- butene |
| 2 | 1-Chlorobutane on reaction with alcohols potash gives. | A. 1- butane<br>B. 1-butanol<br>C. 2- butane<br>D. 2- butanol |
| 3 | The product obtained on heating n-heptane with Cr2O3 ____Al2O3 at 600 $^{o}$C is. | A. Cycloheptane<br>B. Methyl cyclohexane<br>C. Benzene<br>D. Teluene |
| 4 | Bromination of n-butane produces. | A. I-bromobutane as the major product<br>B. 2- bromobutane as the major product<br>C. Both I - bromo and 2-bromobutane with equal percentage<br>D. Both i-bromo and 2-bromo products whose percentage depends upon temperature. |
| 5 | Each of the following compound react with Grignard's reagent to form alkane exxcept. | A. Ethanal<br>B. Ethanoic acid<br>C. Ethanol<br>D. Ethync |
| 6 | In the reaction RCO2Na + Na OH (CaO) _____ RH, we eliminate carboxylate group as. | A. CO2<br>B. Na2CO3<br>C. -CO<br>D. CaCO3 |
| 7 | During the preparation of ethane by Kolbe's electrolytic method using inert electrodes the pH of the electrolyte. | A. Increases progressively as the reaction proceeds<br>B. Decreases progressively as the reaction proceeds<br>C. Remains constant throughout the reaction<br>D. May decrease if the concentration of the electrolyte is not very high |
| 8 | Which of the following hydrocarbon cannot be obtained on reacting chloomethane with sodium metal in the presence of dry ether. | A. C4 H10<br>B. C2H6<br>C. C2H4<br>D. C3H8 |
| 9 | A mixture of ethyl iodide and n-propyl iodide is subjected to Wurts reaction. The hydrocarbon that will nto be formed is | A. n-butane<br>B. n-propane<br>C. n-pentane<br>D. n-hexane |
| 10 | The IUPAC name of HCOOCH3 is. | A. Methoxy methanol<br>B. Ethanoic acid<br>C. Methyl methanoate<br>D. Methoxy methane |
| 11 | The IUPAC name of HOCCH2CH2CH2COOH is | A. 4- formylbutanoic acid<br>B. 5- formylpentanoic acid<br>C. 4- carboxybutanal<br>D. 5- carboxypentanal |
| 12 | The IUPAC name of C2(CN)3 is | A. 2,3-dicvano butanedinitrile<br>B. 2,3 -dicyano -2- butenedinitrile<br>C. 1,1,2,2-tetrcyanoethane<br>D. 1,1,2,2, tetracyanocthenc |
| 13 | IUPAC name of HCONH2 is. | A. Methanamide<br>B. Methanoylamine<br>C. Ammoethanal<br>D. Formanide |
| | | A. Prepanoic anhydride |

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| 14 | The IUPAC name of C2H3, CO , OC OC2H5 in | B. Ethanoic anhydride<br>C. Diketoethoxy ether<br>D. None of the above |
| 15 | The compound (CH3)3COH according to IUPAC is known as. | A. Tert Butanol<br>B. 2,2 -Dumethyl -Propanol<br>C. 2- Methyl -2-propanol<br>D. Tert Alcohol |
| 16 | The IUPAC name of ethylene oxide is. | A. Epoxy methane<br>B. Oxcethene<br>C. Methoxymethane<br>D. All of the above |
| 17 | The IUPAC suffix used for _____ NC group is | A. Cyanide<br>B. Isocyanides<br>C. Carbylamines<br>D. Nitrite |
| 18 | When FeSO4 is added in the sodium extract the compound formed is. | A. Only Na4[Fe (CN)6]<br>B. Only Fe (OH)2<br>C. Only Na2So4<br>D. Mixture of all these |
| 19 | Presence of nitrogen in organic compound to tested as. | A. Nitrogen gas<br>B. NH3<br>C. NO<br>D. Amide |
| 20 | Essential oils are purified by which of the following methods. | A. Steam distillation<br>B. Sublimation<br>C. Crystallization<br>D. Fractional crystallization |
| 21 | The most suitable method of separation in mixture of o-and p- nitrophenol is. | A. Steam distillation<br>B. Chromatography<br>C. Ion-exchange<br>D. Sublimation |
| 22 | Beillstein test is used for. | A. Cl<br>B. N2<br>C. CO2<br>D. Na |
| 23 | The simplest formula of a compound containing 50% of element X | A. XY2<br>B. XY<br>C. X2Y<br>D. None of the abvoe |
| 24 | The molar mass of an organic acids is determined by | A. Depression of freezing point<br>B. Elevation of boiling point<br>C. Volumetric method<br>D. Victor Myer's method |
| 25 | Phosphorus is detected by fusing the organic compound with -------followed by extraction with H2O | A. HNO3<br>B. H2SO4<br>C. Sodium per oxide<br>D. Ozone |
| 26 | Carbon and Hydrogen are estimated by | A. Liebig's method<br>B. Kjeldhal's method<br>C. Carries method<br>D. None of the above |
| 27 | In Dumas method, the volume of the gas collected is equivalent to which of the following gases set free from the compound. | A. Ammonia<br>B. O2<br>C. N2<br>D. NO |
| 28 | The stationary and mobile phases in paper chromatography are. | A. <sup>Liquid/Liquid</sup><br>B. Solid /Liquid<br>C. Liquid/Solid<br>D. Gas/solid |
| 29 | Two solids A and B have appreciable different solubility in water but their m.p. are very close. The mixture A and B can be separated by. | A. Sublimation<br>B. Distillation<br>C. Fractional crystallization<br>D. Specific rotation |
| 30 | The function of boiling the sodium extract with conc. HNO3 before testing the halogens is | A. To make solution clear<br>B. To make the solution acidic<br>C. To bring common ion effect<br>D. To destroy CN- and S2- ion |