PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Which of the following statement is not correct.	A. The element with highest IE belongs to group 18 B. In each period the element with lowest IE belongs to group I C. In each period the element with highest IE is a noble gas D. In the second period as we move from left to right, ionization energy increases regularly.
2	With which one of the following configurations, the lowest value of first IE is associated.	A. ls2, 2s2, 2p6, 3s1 B. 1s2, 2s2, 2p5 C. 1s2, 2s2, 2p6 D. 1s2, 2s2, 2p6, 3s2, 3p2
3	Which of the following configuration is associated with biggest jump between second and third IE.	A. 1s2, 2s2, 2p2 B. 1s2, 2s2, 2p6, 3s1 C. 1s2,2s2,2p6, 3s2 D. 1s2, 2s2,2p6
4	The element with the highest first ionization potential is.	A. Boron B. Carbon C. Nitrogen D. Oxygen
5	The correct order of second ionization potential of carbon nitrogen, oxygen and fluorine is.	A. C > N > O > F B. O > F > N > C C. O > N > F > C D. F > O > N > C
6	The first ionization energy in electron volts of nitrogen and oxygen atoms are respectively given by.	A. 14.6, 13.6 B. 13.6, 14.6 C. 13.6, 13.6 D. 14.6, 14.6
7	The process requirieng the absorption of energy of.	A. F = F B. Cl = Cl C. H = H D. O = O
8	The ionization energy of N is more than that of oxygen becasue.	A. Nitrogen has half filled p orbitals B. Nitrogen atom is smaller in size than oxygen atom C. Nitrogen contains less number of electrons D. Nitrogen is less electronegative
9	Which of the following iso -electronic ion would require least energy for the removal of electron.	A. Ca2+ B. Cl- C. Ar- D. K+
10	Which of the following iso electronic spices has the highest IE.	A. Ne B. Na+ C. F D. O2-
11	The ionization potential of K would be numerically equal to.	A. Electron affinity of Ar B. Electromagnetically of K C. Electron affinity of K+ D. ionization energy of Ca
12	Which ionization Potential in the following equations involves the greatest amount of energy.	A. Na = Na+ + e B. K = K+ +e C. C2+ = C3+ +e D. Ca+ = Ca2+ + e
13	The decreasing order of the second ionization energies of K, Ca and Ba is	A. K > Ca > Ba B. Ca > Ba > K C. Ba > K > Ca D. K > Ba > Ca
14	Which of the following elements would have the lowest first ionization energy	A. Mg B. Rb C. Li

		D. Ca
15	In group 17, the element with highest first ionization enthalpy belongs to.	A. Period 1 B. Period 2 C. Period 7 D. Period 6
16	In whihc period, the element with least ionization enthalpy belong to	A. Group I B. Group 2 C. Group 17 D. Group 18
17	Which of the following represents the correct order of ionic radii	A. La+ < Na+ < K+ < Rb+ B. Li+ > Na + > K+ >Rb+ C. Li = Na + = K+ = Rb+ D. Rb +> Na+> K+ > Li+
18	Which of the following ions does not have the electronic configuration same as that of neon.	A. F- B. O2- C. Na+ D. Ca2+
19	The size of iso electronic species - F- , Ne, and Na+ is affected by	A. Nuclear charge (Z) B. Valence principal quantum number (n) C. Electron electron interaction in the outer orbital D. None of the factors because their size to the same.
20	Which of the following ions is smallest in size.	A. F- B. Cl- C. I- D. Br-
21	The ions Sc3+ , Ca2+ and K+ have same electronic configuration as that of.	A. Neon B. Argon C. Krypton D. Xenon
22	Atomic volume of C, N, O and F are in the order	A. C > N > F > O B. C > N > O > F C. F > O > N >C D. N > C > O >F
23	The correct order of ionic radii for the following ions is.	A. S²- < P³⁻ < Cl- < K+ B. Cl- > S2+ > P3- > k+ C. K+ > Cr > S2+ > P3- D. P3+ > S2 > Cl - > K+
24	In graph of atomic volume versus atomic weight the elements corresponding to peaks in the curve belong to.	A. Group 1 B. Group 18 C. Group 4 D. Group 14
25	An element with atomic number 20 is placed in which period of the periodic table.	A. 1 B. 2 C. 3 D. 4
26	Which group contains elements that exist as monoatomic molecules.	A. 1 B. 2 C. 14 D. 18
27	Elements in the same vertical group of the periodical have same	A. Number of electron B. Atomic number C. Number of valence elections D. Electronic configuration
28	In the long form of periodic table, elements are arranged according to.	A. Increasing atomic number B. Decreasing atomic number C. Increasing atomic mass D. Decreasing atomic mass
29	Lother Meyer plotted a graph showing variation of.	A. Atomic volume with increase in atomic number B. Atomic volume with increase in atomic weight C. Atomic redii with increase in atomic weight. D. Atomic weight which increase in atomic number
30	The law of trinds was proposed by	A. Dobereiner B. Newlands C. Lother Mayer D. Chancourtois