PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Anything that influence the valence electrons will affect the chemistry of the element Which of the following factors does not affect the valency shell.	A. Valence principle quantum number in B. Nuclear charge (Z) C. Nuclear mass D. Number of core electrons
2	An element having low IE and low EA is likely to belong to.	A. Group IA B. Group IB C. Group VII A D. Group VIII
3	The element Uuu has atomic numebr	A. 102 B. 111 C. 101 D. 110
4	Which of the following pairs shows diagonal relationship	A. Li and Mg B. Na and K C. Zn and Cd D. Li and BE
5	A property which gradually increases on moving down group in the periodic table is	A. Ionization enthalpy B. Elemtronegativity C. Electron affinity D. atomic size
6	The variable valency is generally observed in case of.	A. Transition elements B. Inert gases C. Normal elements D. Non- metallic elements
7	The element with atomic numebr greater than 100 are known as	A. Trans uranium elements B. Trans fermium elements C. Actinides D. Lanthanides
8	The electromagevitiy of the following elements increase in the order	A. F > Cl > O > S B. S > Cl > O > F C. F > O > N > C D. C > O > N > F
9	Electronegativity is given by	A. Average of first and second ionization energies. B. Average of first and second electron affinites C. Average of ionization energy and electron affinity D. None of the above
10	Keeping in view the periodic law and periodic table, suggest which of the following elements should have maximum electronegative character.	A. Oxygen B. Nitrogen C. Fluorine D. Astatine
11	The electronegativity of the following elements increases in the order.	A. C,N, Si, P B. N, Si, C,P C. Si, P, C, N D. P, Si, N, C
12	The most electronegative element of the third period is.	A. F B. P C. Br D. Cl
13	An element with high electronegativity has	A. High IE and high EA B. High IE and low EA C. Low IE and High EA D. Low IE and low EA
14	In each period the most electro negative element belongs to.	A. ^{Group -1} B. Group -17 C. Group -2 D. Group -18

15	The element having highest ionization energy and least electron affinity belong to	A. Period 1 , group 18 B. Period 2, group 17 C. Period 2, group 1 D. Period 2, group 2
16	The electron gain enthalpy of chlorine is -349 Kj mol-1 ionization energy of Cl would be.	A. -349 kJ mol -1 B. 349 kJ mol-1 C. -698 kJ mol-1 D. 698 kJ mol-1
17	The electronic configuration of some elements are given below. The element with highest electron affinity is	A. 1s2, 2s2, 2p3 B. 1s2, 2s2, 2p4 C. 1s2, 2s2, 2p5 D. 1s2, 2s2, 2p6
18	In each period the element with lest electron affinity belongs to.	A. Group 1 B. Group 14 C. Group 17 D. Group 18
19	The correct order of second ionization potential of carbon , nitrogen, oxygen and fluorine is.	A. C > N > O > F B. O > N > F > C C. O > F > N > C D. F > O > N > C
20	Electron affinities of halogens are in the order.	A. F > Cl > Br > I B. Cl > F > Br > I C. Cl > Br > I > F D. Cl > Br > F > I
21	The correct order of electron affinities is.	A. C > Si, > Na > Ar B. Si > Cl > Na > Ar C. C > Na > Si > Ar D. C > Si > Ar > Na
22	The correct order of electron affinities of Sl, P, and Cl is.	A. P > Si > Cl B. Cl > P > Si C. Cl > Si, > P D. Si > P . Cl
23	The magnitude of electron affinity depends on.	A. Atomic size B. Nuclear charge C. Electronic configuration D. All of the above
24	The elements with highest electron affinity belongs to.	A. Period 2 , group 17 B. Period 3, group 17 C. Period 2, group 18 D. Period, 2 , group 1
25	In a period, the element with biggest electron affinity belong to.	A. Group 1 B. Group 2 C. Group 17 D. Group 18
26	The first ionization energy of Mg is lower than	A. Na B. Ca C. Al D. Be
27	Which of the following elements has the highest ionization energy.	A. Na B. Si C. Ar D. Cl
28	Which one of the following statement is incorrect in relation to ionization enthalpy.	A. Ionization enthalpy increase for each successive electron B. The greatest increase in ionization enthaly is experienced on removal of electron from core noble gas configuration C. End of the valence electron is marked by a big jump in ionization enthalpy D. Removal of electron from orbitals bearing lower value to easier than from orbital having highest n value.
29	Which of the following elements has the highest third ionization energy.	A. Sodium B. Magnesium C. Aluminum D. Silicon
30	Among the elements A,B,C and D having atomic numbers 9,10,11, and 12 respectively, the correct order of ionization energies is.	A. A > B > C >D B. B > A > D > C C. B > A > C > D D. D > C > B > A