PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Which compound among the following does not contain an ionic bond.	A. NaOH B. HCl C. KaS D. LiH
2	Which of the following compounds is electrovalent in nature.	A. SO2 B. ICl C. KBr D. CHl3
3	Which of the following compound does not following octet rule.	A. CS2 B. PBr3 C. IBr D. Br F3
4	An ionic compound X + Y - is most likely to be formed if	A. Ionization enthalpy of X is high electron gain ethalpy of Y is low B. Ionization enthalpy of X is high electron gain enthalpy of Y is high C. Ionization enthalpy of X is low,electron gain enthalpy of Y is low D. Ionization enthalpy of X is low electron gain enthalpy of Y is high
5	Four elements A, B,C,D have atomic numbers Z , 1 ,Z, Z +1 and Z +2 Respectively If Z is 9, then bond between which pair of elements will be ionic.	A. A and C B. D and C C. D and B D. B and C
6	Which of the following orbitals has maximum penetration effect.	A. s B. p C. d D. f
7	Among the elements of third period, the element with lowest boiling point belongs to group.	A. 1 B. 14 C. 16 D. 18
8	Among the elements of second period the element with highest melting point belongs to group.	A. 1 B. 14 C. 17 D. 18
9	Which of the following has the highest melting poing.	A. NaCl B. Li Cl C. KCl D. Rb Cl
10	Considering the elements F, Cl ,O and N, the correct order of their chemical reactivity in terms of oxidizing property is.	A. F > Cl > N B. F > O > Cl > N C. Cl > F > O > N D. O > F > N > Cl
11	Which of the following hydroxides has the maximum solubility in water.	A. Mg (OH)2 B. Ca (OH)2 C. Sr (OH)2 D. Ba (OH)2
12	Which of the following is most soluble in water.	A. BaSO4 B. Sr SO4 C. CaSO4 D. MgSO4
13	Which of the following carbonates decomposes at the highest temperature.	A. Mg CO3 B. CaCO3 C. Sr CO3 D. Ba CO3
14	Which of the following has the greatest metallic character.	A. Na B. Mg C. Al D. Si

15	In each period, the most electropositive element belongs to group.	A. 18 B. 17 C. 1 D. 2
16	The common oxidation state of lanthanides is.	A. +3 B. +2 C. +1 D. +4
17	Which of the following elements forms maximum number of compounds.	A. Carbon B. Silicon C. Hydrogen D. Fluorine
18	Which of the following elements is most electropositive.	A. C B. N C. O D. Be
19	The most electronegative and the most electromotive elements of the first period is	A. H and He B. Na and Cl C. Li and F D. H and He
20	Among the elements A, B,C and D having atomic numbers 7,8,9 AND 12 Respectively, the element with smallest size and highest IE is.	A. A B. B C. C D. D
21	Which one of the following sets of elements has the strongest tendency to form positive ions in gaseous state.	A. Li, Na, K B. F, Cl, Br C. Be, Mg, Ca D. O, S , Se
22	Which one of the following sets of elements has the strongest tendency to form negative ions in gaseous state.	A. NA, Mg, Al B. Ca, V, Cr C. N , O, F D. Ga, In, tl
23	Which of the following statements is incorrect.	A. The elements of group 18 are known as aerogens. B. Group 2 elements are all metals C. Metallic character increase on going down a group D. All the elements belonging to a particular period have same valence shell configuration.
24	Which one of the following pairs are chemically dissimilar.	A. Na and K B. Ba and Sr C. Zr and HF D. Ca and Zn
25	Among group IA elements, melting point	A. Increases down the group B. Decreases down the group C. Do not show any regular trend D. Remains constant
26	Considering the element B,C, N, G and Si, the correct order of their non- metallic character is.	A. B > C > Si > N > F B. Si > C > B > N > F C. F > N > C > B > Si D. F > N c > > SI > B
27	Considering the elements B, Al, Mg and K , the correct order of their metallic character is.	A. B > Al > Mg > K B. Al > Mg > B > K C. Mg > Al > K > B D. K > Mg > Al > B
28	Which of the following generally increases on going from top to bottom in a group.	A. Metallic charcter B. Electronegativity C. Oxidising behaviour D. Raducing behaviour
29	Berllium has diagonal relationship with	A. Li B. Al C. B D. Na
30	A trend which is common to elements of both the group IA and group VII A ongoing from top to bottom.	A. Boiling point increases B. Electron affinity increases C. Oxidizing power increases D. Ionization energy decrease