

PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Mostly used solvents for ionic compounds.	A. Liquid ammonia B. Liquid SO ₂ C. Liquid HF D. All above
2	By applying an external force the ionic solid can be easily broken to powder form so the ionic solid are highly	A. Hard B. Brittle C. Tough D. Soft
3	The unit of sodium chloride structure is.	A. Linear B. Cubic C. Tetrahedral D. Square planar
4	Solid substances consist of an ordered array of ions and solid as a whole is electrically.	A. Conductor B. Neutral C. Acidic D. Basic
5	Variable electrovalency is due to the following reasons.	A. Unstable configuration of core B. Inert electron pair effect C. All of above D. None of above
6	Stable metal ions structures are.	A. Noble gas structure B. Inert electron group structure C. Transition metal in structure D. All of the above
7	The bond formed by complete transfer of electrons from electropositive to more electronegative atom is called.	A. Ionic bond B. Covalent bond C. Metallic bond D. Coordinate bond
8	Non localised bonds are referred as	A. Metallic bond B. Long range bonds C. Ionic bond D. Covalent bonds
9	Ionic bond are also forces called as.	A. Polar bond B. Electrovalent bond C. Non polar bond D. Both A and B
10	When two H atoms approach each other then forces operate.	A. Attractive forces B. Repulsive forces C. Attractive and repulsive D. None of above
11	A stable molecule is a group of atoms held together by	A. Chemical forces B. Physical forces C. Valence force D. None of above
12	The forces which hold the atoms together in a molecule is called	A. Ionic bond B. Covalent bond C. Coordinate bond D. Chemical bond
13	A molecule MX ₄ has a square planar shape, The number of non bonding pairs of electrons around M is .	A. 2 B. 1 C. 0 D. 3
14	Which of the following molecule does not contain the covalent bond between similar atoms.	A. N ₂ H ₄ B. F ₂ O ₂ C. H ₂ F ₂ D. H ₂ O ₂
15	In the formation of H ₂ O molecule, the oxygen atom makes use of.	A. 2p orbitals B. sp hybrid orbitals C. Sp ² hybrid orbitals D. Sp ³ hybrid orbitals

16	The geometry of the molecule is primarily decided by	A. Bond pairs around the central atom B. No of k bond around the central atom C. No of bond pairs as well as lone pairs around the central atom D. No. of lone pairs on central atom
17	The configuration of valence shell of certain atom X is $3s^2, 3p^5$, which valences can it exhibit.	A. 1,3 only B. 1,5 only C. 1,3,5,7 D. 1,3,4
18	In vinyl cyanide, the number of a bonds in	A. 2 B. 3 C. 1 D. 4
19	In which of the following species the bonds are non directional.	A. NCl_3 B. RbCl C. BeCl_2 D. BCl_3
20	Which of the following statements is correct.	A. A sigma bond is weaker than a pi bond B. There are four coordinate bonds in the Lewis structure of NH_4^+ ion. C. The 1 covalent bond is directional in nature D. A single bond between the two atoms cannot be re bond.
21	The Lewis structure of which of the following does not have coordinate bond.	A. SO_2 B. HNO_3 C. H_2SO_4 D. HNO_2
22	Which of the following molecule contains two dative bonds according to Lewis structure.	A. NH_3 B. SO_3 C. PCl_5 D. BF_3
23	The number of electrons involved in bonding in Lewis structure of oxalate ion is	A. 20 B. 14 C. 22 D. 18
24	In the Lewis structure of H_2SO_4 molecule the total number of unshared electrons in valence shell of various atoms is.	A. 8 B. 16 C. 12 D. 20
25	Which of the following statements in incorrect.	A. Sodium hydride is ionic B. Beryllium chloride is covalent C. CCl_4 gives a white ppt with AgNO_3 solutions. D. Bonds in NaCl are non directional
26	Which of the following is not a characteristic of covalent compound.	A. They have low melting and boiling points. B. They ionize on dissolution in polar solvents C. Their molecules have definite geometry D. They are generally insoluble in water
27	In bi sulphate ion, the formal charge on sulphru atom is.	A. +1 B. +2 C. +4 D. +6
28	Which of the following element has six electrons in the valance shell but cannot exhibit a maximum co valency of six.	A. Sulphur B. Oxygen C. Salenium D. Both A and B
29	Which element out of the following can exhibit a maximum co valency of seven.	A. Chlorine B. Fluorine C. Sulphur D. Both Cl and F
30	The Lewis formula of SOCl_2 , the total number of bond pairs and lone pairs of electron around sulphur are.	A. 2,1 B. 2,2 C. 3,1 D. 3,0

