PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Metals are	A. Transparent B. Tranalucant C. Opaque D. None of above
2	Metal are generally elements	A. Electronegative B. Electropositive C. Neutral D. None of the above
3	Metal crystallize is system having co ordination number	A. 8 B. 12 C. 14 D. any one of above
4	H-Bonding also ox in ling system like	A. Protein B. DNA C. Botha A and B D. None of above
5	H-Bond has a preferred bonding direction like	A. Ionic bond B. Covalent bond C. Co ordinate bond D. None of these
6	H-Bond has more energy than the van der Waals forces i.e.	A. 1.0 kcal/mole B. 2.0 k cal/mole C. 10.0 kcal/mole D. 20. 0 kcal mole
7	Excluding H-atom , Hydrogen bond never involves more than atoms.	A. One B. Two C. Three D. Four
8	Hydrogen bond is not electrostatic in nature is stated by	A. Electrostatic approach B. Valence bond approach C. Molecular orbital approach D. None of the abvoe
9	In order to understand the nature of H , bond the theory has been suggested.	A. Electrostatic approach B. Molecular orbital approach C. Valance bond approach D. All the above approaches
10	Example of inter molecular H-bonding is	A. NH3 and H2O B. HF C. CH3COOH D. All of abvoe
11	Example of intra molecular hydrogen bonding.	A. O-nitrophenol B. O-hydroxy benzaldehyde C. O- hydroxy benzoic acid D. All of the above
12	The bond order for BO molecule is.	A. 2.5 B. 3.0 C. 2.0 D. 3.5
13	The bond order gives the following valuable information.	A. Stability of the molecules of ions B. Bond dissociation energy and bond length C. Magnetic properties D. All of the above
14	d2 sp3 is oriented in a manner	A. Trigonal B. Tetrahedral C. Octahedral D. Trigonal bipyramidal
15	PCl5 is an example of hybridization	A. d sp³ B. d2 sp2 C. sp2

		D. sp3
16	The bond along Sp2 hybridization is.	A. 180^o B. 120^o C. 109.5^o D. 160^o
17	BCl3 is an example of hybridization	A. sp B. sp2 C. sp3 D. None of above
18	A covalent bond which is formed between two atoms by the overlap of atomic orbitals along their axis is called.	A. Pi bond B. Sigma bond C. Polar bond D. Non polar bond
19	Valence bond theory is also called as	A. Electron pair theory B. Band theory C. Electron gas theory D. Electron pool theory
20	Pauling has suggested that the calculate of energy can be improved by considering.	A. Screening effect B. Polarization effect C. Both A and B D. None of abvoe
21	Molecules have zero dipole moment	A. CO2 B. BCl3 C. CH4 & CCl4 D. All above
22	The polarity of bonds can lead to polarity of molecules and affect	A. Melting point B. Boiling point C. Solubility D. All of above
23	The bond length is measured by	A. X-ray diffraction B. Neutron diffraction C. Microwave spectroscopy D. All of above
24	The types of coordinate compounds.	A. Labila B. Inert C. Both A and B D. None of above
25	Co ordinate compounds are	A. Polar B. Non polar C. Dem polar D. None of above
26	Al2Cl6 is an example of	A. Ionic bond B. Covalent bond C. Co ordinate bond D. Metallic bond
27	The acetylene molecule contain a	A. Single bond B. Double bond C. Triple bond D. Co ordinate bond
28	A type of a chemical bond which is formed by the mutual sharing of electrons between combining atoms of the same or different elements is called.	A. Ionic bond B. Covalent bond C. Co ordinate Covalent bond D. Metallic bond
29	Titanium dioxide shows the lattice strcuture.	A. Flluorite B. Rutile C. Wurtzite D. Zeolite
30	The coordination number of closely packed hexagonal is.	A. 4 B. 6 C. 8 D. 12