PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	The pH of the tears is	A. 7.0 B. 7.4 C. 7.8 D. 8.2
2	The concentration of OH- ions in a certain household ammonia solution is 0.0025. This ammonia solution is.	A. Basic B. Acidic C. Neutral D. None of above
3	The value of Kw increase with temperature because the ionization of water.	A. Positive B. Negative C. Endothermic D. Exothermic
4	The ionic product equilibrium constant is.	A. Ka B. Kb C. Kc D. Kw
5	pH of pure water at 25 ^{o}C. kw = 1 x 10^{-4}	A. 0 B. 7 C. 14 D. None of above
6	A chemical reaction resulting in a change in the electric charge on the reacting particles may be called as.	A. Add ion reaction B. Redox reaction C. Elimination reaction D. Chain reaction
7	Relative order of acidity of oxy acid	A. HClO > HClO2 > HClO3 > HClO4 B. HClO4 > HClO3 >HClO2 > HClO C. HClO3 > HClO2 > HClO > HClO4 D. HClO2 > HClO4 > HClO3 > HClO
8	Relative order of acidity of HF, HCl, HBr, and HI acids is	A. HCl > HBr > HI > HF B. HF > HCl > HBr > HI C. HI > HBr > HCl >HF D. HF > HI > HCl > HBr
9	All the strong acids have very close pKas value and they appear to have nearly equal strengths in aqueous solutions. The phenomenon is called as.	A. Levelling effect B. Differnetiating effect C. Levelling solvent D. Differnetiating solvent
10	The one in which the acceptor atom is of low positive charge, Large size and has several outer electrons which can be easily excited is a.	A. Soft base B. Hard Base C. Soft acid D. Hard acid
11	According to SHAH concept the Lewis bases were classified on the basis os.	A. Charge ion size B. Polarization consideration C. Electron and co coordinating ability D. All of above
12	According to SHAB, Lewis acid are divided into.	A. Two classes B. Three classes C. Four classes D. None of above
13	According to Usanovich concept a base is defined as any species.	A. Capable of giving up anions B. Combining with cations C. Neutralizing an acid to give a salt D. All of above
14	Lux -Flood concept is a dono-acceptor system of.	A. Proton B. Electron pair C. Neurtron D. Oxide ion

15	Lewis concept explain the formation of	A. Ionic bond B. Covalent bond C. Co-ordinate bond D. Chemical bond
16	Bases and reducing agents are electron giving agents and also called as.	A. Electrodotic B. Electrophile C. Nucleophile D. None of above
17	The concept is also known as electron pair donor acceptor system.	A. Bronsted Lowery B. Lewis C. Lux -Flood D. Usanovich
18	The concept is also known as proton donor acceptor system.	A. Bronsted Lowery B. Lewis C. Lux Flood D. Usanovich
19	Arrhenius concept explained	A. Constant heat of neutralization B. Quantitative determination of acid base strength C. Catalytic property of acid D. All above
20	"Acids are substance whose aqueous solutions turned blue litmus red and tasted sour" stated by	A. Davy B. Liebig C. Boyle D. Rouelle
21	How many sigma and pi bonds are there in a CO2 molecule.	A. 2 sigma B. 2 sigma and 4 pi C. 2 sigma and 2 pi D. 4 sigma and no pi
22	Which of the following interaction is the strong.	A. Dipole -dipole B. Ion induced dipole C. Ion -dipole D. Dipole induced diple
23	Covalent compound are soluble in	A. Polar solvents B. Non polar solvent C. Concentrated acids D. All solvent
24	The maximum covalence of an element equal to.	A. The number of unpaired d electrons B. The number of paired p electrons C. The number of unpaired a and P electors D. The actual number of a and P electrons in the outermost shell
25	Arrange the following in order of increasing boiling point.	A. CH3OH < CH3Cl < RbCl < CH4 B. CHOH < CH4 < CH3Cl < RbCl C. RbCl < CH3Cl < CH3OH < CH4 D. CH4 < CH3Cl < CH3OH < RbCl
26	Which of the following statements is worng.	A. Covalent compounds are generally soluble is polar solvents B. Covalent compounds have low melting and boiling point. C. Ionic solids do not conduct electricity is solid state D. Ionic compounds conduct electricity in the fused state.
27	Which of the following statements is wrong.	A. Covenant compounds are generally soluble is polar solvents. B. Covalent compounds have low melting and boiling points C. Lower than that of separate H atoms D. Sometimes lower and sometimes higher than that of separate H
28	Which two atoms of hydrogen combine to form a molecule of hydrogen gas. the energy of the hydrogen molecule is.	A. Higher than that of separate H atoms B. Equal to that of separate H atoms C. Lower than that of separate H atoms D. Sometimes lower and sometimes higher than that of separate H.

		higher than that of separate H_2
29	The boiling point of water is unexpectedly high because.	A. H_2O molecule in linear B. Sp3 hydrogen bonding is involved in the formation of water C. There is hydrogen bonding and consequent association of H_2O molecules. D. Oxygen is the first member of the VI group
30	Which of the following statements regarding covalent bond is false.	A. The electrons are shared between atoms. B. The bond in non -directional C. The strength of the bond depend upon the extent of overlapping D. The bond formed may be polar or non-polar