

PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Commercial or the phosphoric acid is pure.	A. 37.0% B. 82.98% C. 88.25%
2	The orientation of a crystalline surface is confidently defined in terms of.	D. 90.12% A. Lijima Indices B. Miller indices C. Clausen indices D. None
3	The correct order of thermal stabilities of hydriedes of group 15 is.	A. NH3 > PH3 > AsH3 > BiH3> SbH3 B. NH3 > PH3 >AsH3 > SbH3> BiH3 C. NH3 < PH3 < SbH3 > AsH3 > BiH3 D. BiH3 > Sb H3 > AsH3 > PH3 > NH3
4	Which of the following techniques is capable of seperating minute quantities of the substances in a relatively short times with high resolutions.	A. Gel electrophoresis B. Capillary electrophoresis C. GC D. HPLC
5	The correct order of ionic radii for the following ions is.	A. S ² - ⁢ P ^{3- } ⁢ CI- ⁢ K+ B. CI- > S2+ > P3- > k+ C. K+ > Cr > S2+ > P3- D. P3+ > S2 > CI - > K+
6	Glass electrode cannot be used to measure the pH of pure.	A. Acetic acid B. Ethyl alcohol C. Gelatin D. All above
7	Bases and reducing agents are electron giving agents and also called as.	A. Electrodotic B. Electrophile C. Nucleophile D. None of above
8	The increase in boiling points of noble gases from He to Xe is due to the	A. Decreases in ionization energy B. Increases in polarizability C. Increase in electron affinity D. Increase in atomic volume
9	In which of the following characteristics does hydrogen resemble halogens.	A. Hydrogen is the lightest gas B. H atoms contains one electron each C. Hydrogen forms ionic hydrides with alkali metals D. Hydrogen has three isotopes.
10	According to Fajns rules, which one of following results in increased ionic nature of the covalent bond.	A. Larger cation and smaller charges on anion B. Larger cation and larger charge on anion C. Smaller cation and smaller charge on anion D. Smaller cation and larger charge on anion
11	The valence shell electronic configuration of group III A is.	A. ns1 p2 B. ns2 p1 C. ns3p2 D. ns2p2
12	HCIO2 gives the structure of a.	A. Linear B. Angular C. trigonal pyramidal D. Tetra hddral
13	When two bodies have equality of temperature with a 3rd body they in turn have equality of temperature with each other. This is a statement of.	A. First law of thermodynamics B. Zeroth law of thermodynamics C. Second law of thermodynamics D. Third law of thermodynamics

14	The IUPAC name of HCOOCH3 is.	A. Methoxy methanol B. Ethanoic acid C. Methyl methanoate D. Methoxy methane
15	Process of separating the racemic mixture into optically active isomers is known as.	A. Resolution B. Racemisation C. Walden inversion D. Epimerization
16	Potassium crystallizes in a body centered lattice. Hence, the coordination number of potassium in potassium metal is.	A. 4 B. 6 C. 8 D. 12
17	Which is major component of Bordeaux mixture.	A. Copper sulphate <0:p>>/o:p> B. Sodium chloride<0:p> C. Sodium chloride<0:p> C. Calcium chloride<0:p> D. Magnesium sulphate<0:p>
18	Pick out the incorrect statement for Xe F6	A. XeF6 is hydrolyzed practically to form XeOF4 B. It reacts with SiO2 to form Xe F4 C. On complete hydrolysis, it forms XeO3 D. It acts as F acceptor when treate with alkali metal fluoride, but canno act as F donor to form complexes.
19	What is the raw material of sugar industry.	A. Sugar cane B. Potato C. Carrot D. Sugar heat E. Both A and C
20	Which of the following is not a proper use of Ni.	A. It is used as catalyst B. It is used in alloy formation C. It is used in the preparation of Monel metal D. It is attached by alkalis