

## NAT II Biological Science Chemistry

Sr	Questions	Answers Choice
1	For most of the chemical reactions the rate of reaction	A. Increase as the reaction proceeds B. Decrease as the reaction proceeds C. May increase or decrease during the reaction D. Remains constant as the reaction proceeds
2	The unit rate of a reaction can be increased in general by all the factors except by	A. Using a catalyst B. Increasing temperature C. Increasing the activation energy D. Increasing the conc. of reactants
3	When $\text{KClO}_3$ is heated, it decomposes into $\text{KCl}$ and $\text{O}_2$ . If some $\text{MnO}_2$ is added, the reaction goes much faster because	A. $\text{MnO}_2$ decomposes to give $\text{O}_2$ B. $\text{MnO}_2$ provide heat by reacting C. Better contact is provided by $\text{MnO}_2$ D. $\text{MnO}_2$ acts as a catalyst.
4	The rate of reaction between A and B increases by a factor of 100, and when the concentration with respect to A is increased 10 folds, the order of reaction w.r.t. A is	A. 10 B. 1 C. 4 D. 2
5	A certain current liberate 0.5 g of hydrogen in 2 h. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphate solution?	A. 12.7 gm B. 15.9 gm C. 31.8 gm D. 63.5 gm
6	A current of 9.65 ampere flowing for 10 minutes deposits 3.0 g of the metal which is monovalent. The atomic mass of the metal is	A. 10 B. 50 C. 30 D. 96.5
7	A solution of sodium sulphate was electrolysed using some inert electrodes. The products at the electrodes are	A. $\text{O}_2$ , $\text{H}_2$ B. $\text{O}_2$ , $\text{Na}$ C. $\text{O}_2$ , $\text{SO}_2$ D. $\text{O}_2$ , $\text{S}_2\text{O}_8^{2-}$
8	A cell constant is generally found by measuring the conductivity of aqueous solution of	A. $\text{BaCl}_2$ B. $\text{KCl}$ C. $\text{NaCl}$ D. $\text{MgCl}_2$
9	An electrolyte	A. Forms complex ions in solution B. Gives ions only when electricity is passed C. Possesses ions even in solid state D. Gives ions only when dissolved in water
10	When electricity is passed through molten $\text{Al}_2\text{O}_3 + \text{Na}_3\text{AlF}_6$ and 13.5 gms of Al are deposited, the number of faraday must be	A. 0.5 B. 1.0 C. 1.5 D. 2.0
11	If a salt bridge is removed electrode is made from which of the following?	A. $\text{ZnCl}_2$ B. $\text{CuSO}_4$ C. $\text{Hg}_2\text{Cl}_2$ D. $\text{HgCl}_2$
12	When quantity of electricity passed is one faraday then the mass deposited at the electrode is equal to	A. One gm. atomic weight B. One gm. equivalent weight C. Electrochemical equivalent D. None of the above
13	Which of the substances Na, Hg, S, Pt and graphite can be used as electrodes in electrolytic cells having aqueous solution?	A. Na, Pt and graphite B. Na and Hg C. Pt and graphite only D. Na and S only

A. Osmotic pressure

14	Which is not a colligative property?	B. Lowering of vapour pressure C. Depression of freezing point D. Elevation of boiling point
15	The molal elevation constant is the ratio of the elevation in boiling point to	A. Molarity B. Molality C. Mole fraction of solute D. Mole fraction of solvent
16	Which inorganic precipitate acts as semipermeable membrane?	A. Calcium sulphate B. Barium oxalate C. Nickel phosphate D. Copper ferrocyanide
17	The movement of solvent molecules through a semipermeable membrane is called	A. Electrolysis B. Electrophoresis C. Osmosis D. Cataphoresis
18	Saturated solution of NaCl on heating becomes	A. Super saturated B. Unsaturated C. Remains saturated D. None
19	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
20	Which of the following is a colligative property?	A. Melting point B. Osmotic pressure C. Freezing point D. Sublimation temperature