

NAT I Medical Chemistry

Sr	Questions	Answers Choice
1	Fluorine does not show positive oxidation states due to the absence of	A. d-orbitals B. s-orbitals C. p-orbitals D. None
2	Saturated solution of NaCl on heating becomes	A. Super saturated B. Unsaturated C. Remains saturated D. None
3	Which of the following units represents largest amount of energy?	A. Calorie B. Joule C. Erg D. Electron vol.
4	Inert pair effect plays an important role in case of	A. F B. Al C. Si D. Tl
5	The dimension of rate constant of a second order reaction involves	A. Neither time nor concentration B. Only time C. Time and concentration D. Time and square of concentration
6	Which of the following species participate in sulphonation of benzene ring?	A. H^+ B. HSO_3^+ C. SO_3 D. SO_2
7	Octane number can be changed by	A. Isomerisation B. Alkylation C. Cyclisation D. All of these
8	Which of the following mineral does not contain Al?	A. Cryolite B. Mica C. Feldspar D. Fluorspar
9	Evaporation of water is	A. An exothermic change B. An endothermic change C. A process where no heat changes occur D. A process accompanied by chemical
10	Which has maximum protein content?	A. Ground nut B. Cow milk C. Egg D. Wheat
11	The effect of increasing the pressure on the following equilibrium $2A + 3B \rightleftharpoons 3A + 2B$ is	A. Forward reaction is favoured B. Backward reaction is favoured C. No effect D. None of the above
12	Bromine is obtained on a commercial scale from	A. Caliche B. Carnallite C. Common salt D. Cryolite
13	Which one is not a pollutant normally?	A. Hydrocarbons B. Carbon dioxide C. Carbon monoxide D. Sulphur dioxide
14	Which of the following has greatest tendency to lose electron?	A. F B. Fr C. S D. Be

15	The number of atoms contained in 11.2 L of SO ₂ at S.T.P are	<p>A. $\frac{3}{2} \times 6.02 \times 10^{23}$</p> <p>B. $2 \times 6.02 \times 10^{23}$</p> <p>C. 6.02×10^{23}</p> <p>D. $4 \times 6.02 \times 10^{23}$</p>
16	Which is true for an element R present in group 13 of the periodic table?	<p>A. It is a gas at room temperature</p> <p>B. It has oxidation state of +4</p> <p>C. It forms R₂O₃</p> <p>D. It forms RX₂</p>
17	A cell constant is generally found by measuring the conductivity of aqueous solution of	<p>A. BaCl₂</p> <p>B. KCl</p> <p>C. NaCl</p> <p>D. MgCl₂</p>
18	When quantity of electricity passed is one faraday then the mass deposited at the electrode is equal to	<p>A. One gm. atomic weight</p> <p>B. One gm. Equivalent weight</p> <p>C. Electrochemical equivalent</p> <p>D. None of the above</p>
19	Which metal is protected by a layer of its own oxide?	<p>A. Al</p> <p>B. Ag</p> <p>C. Au</p> <p>D. Fe</p>
20	Which one is primary alcohol?	<p>A. Buten-2-ol</p> <p>B. Propan-2-ol</p> <p>C. Butane-1-ol</p> <p>D. 2,3-Dimethylhexane-4-ol</p>