

NAT I Medical Chemistry

Sr	Questions	Answers Choice
1	One mole of a gas refers to	A. The number of molecules in one litre of gas B. The number of molecules in one gram of gas C. The number of molecules contained in 12 grams of ^{12}C isotope D. The number of molecules in 22.4 liters of a gas at S.T.P.
2	The kinetic theory of gases predicts that total kinetic energy of a gaseous assembly depends on	A. Pressure of the gas B. Temperature of the gas C. Volume of the gas D. Pressure temperature and volume of the gas
3	Reaction of acids with alcohols is also known as	A. Esterification B. Saponification C. Alkalization D. None
4	An endothermic reaction is one in which	A. Heat is converted into electricity B. Heat is absorbed C. Heat is evolved D. Heat is converted into mechanical work
5	When KClO_3 is heated it decomposes into KCl and O_2 if some MnO_2 is added the reaction goes much faster because	A. MnO_2 decomposes to give O_2 B. MnO_2 provides heat by reacting C. Better contact is provided by MnO_2 D. MnO_2 acts as a catalyst
6	Alum is not used	A. As a mordant in dyeing B. As an insecticide C. In purification of water D. In tanning of leather
7	Bell metal is an alloy of	A. Cu , Zn , and Sn B. Cu , Zn and Ni C. Cu and Zn D. Cu and Sn
8	When pressure is applied to the equilibrium system Ice Water Which of the following phenomenon will happen?	A. More ice will be formed B. Water will evaporate C. More water will be formed D. Equilibrium will not be formed
9	Which of the following process is used to separate insoluble particles from liquids?	A. Separation B. Filtration C. Crystallization D. Condensation
10	Natural fertilizer provides potassium in the form of K_2O (potash)	A. 1.5 kg B. 3 kg C. 4.5 kg D. 6.5 kg
11	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
12	Which one is primary alcohol?	A. Buten-2-ol B. Propan-2-ol C. Butane-1-ol D. 2,3-Dimethylhexane-4-ol

13	A certain liberate 0.5 g of hydrogen in 2 h. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphate solution?	<p>A. 12.7 gm</p> <p>B. 15.9 gm</p> <p>C. 31.8 gm</p> <p>D. 63.5 gm</p>
14	Chile salt petre is	<p>A. NaNO_3</p> <p>B. Na_2SO_4</p> <p>C. KNO_3</p> <p>D. $\text{Na}_2\text{S}_2\text{O}_3$</p>
15	Which of the following is acidic?	<p>A. SO_3</p> <p>B. N_2O</p> <p>C. BeO</p> <p>D. HgO</p>
16	Sodium metal cannot be stored under	<p>A. Benzene</p> <p>B. Kerosene oil</p> <p>C. Alcohol</p> <p>D. Toluene</p>
17	Ethyl chloride on treatment with aqueous alkali gives	<p>A. Ethane</p> <p>B. Ethene</p> <p>C. Ethanal</p> <p>D. Ethanol</p>
18	Which of the following geometry is associated with the compound in which the central atom assumes sp^3 d hybridization?	<p>A. Planar</p> <p>B. Pyramidal</p> <p>C. Angular</p> <p>D. Trigonal bipyramidal</p>
19	If N_A is Avogadro's number then number of valence electrons in 4.2 g of nitride ions N^{3-} is	<p>A. $2.4 N_A$</p> <p>B. $4.2 \times 14.44444465637207 \times 10^{23}$</p> <p>C. $1.6 \times 14.44444465637207 \times 10^{23}$</p> <p>D. $3.2 \times 14.44444465637207 \times 10^{23}$</p>
20	Tollen's reagent is	<p>A. Ammonical cuprous chloride</p> <p>B. Ammonical cuprous oxide</p> <p>C. Ammonical silver bromide</p> <p>D. Ammonical silver nitrate</p>